

Demo of Hydrolysis AlCl_3 , CaCl_2 , Na_2CO_3 , etc.**Hydrolysis:**

- Reaction between a **salt** (ion or ions in a salt) and **water** to produce an **acidic** or **basic** solution.
- Net ionic equations for **hydrolysis**:

An ion + water \rightarrow a molecule or ion + H_3O^+ or OH^-

SPECTATORS- ions which do NOT hydrolyze (need periodic table and acid table to find these)

Spectator Cations (look on per. table)

Group 1 (Alkali Metal ions) eg. Li^+ , Na^+ , K^+ , Rb^+ , Cs^+ , Fr^+
Group 2 (Alkaline Earth ions) eg. Be^{2+} , Mg^{2+} , Ca^{2+} , Ba^{2+} , Sr^{2+} , Ra^{2+}

Spectator Anions (look on acid table)

- Conjugate bases of strong acids.
- **Top 5 ions on the right** side of table.
- ClO_4^- , I^- , Br^- , Cl^- , NO_3^-

(HSO_4^- is not a spectator – it is amphiprotic – will be dealt with later)

- **spectators are eliminated** in net ionic equations (NIE's) for hydrolysis!

Process – if given **salt** (dissociate \rightarrow eliminate \rightarrow evaluate)

1. Write **dissociation** equation
2. Eliminate **spectators**
3. Remaining ions \rightarrow **left** side of table – undergo **acid** hydrolysis – produce H_3O^+
 \rightarrow **right** side of table – undergo **base** hydrolysis – produce OH^-
 \rightarrow **amphiprotic** – determine K_a and K_b to find **dominant** hydrolysis.

Examples:

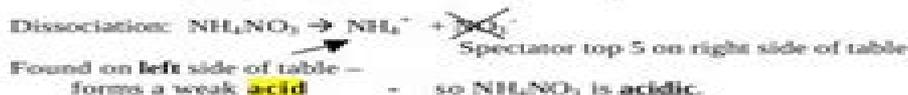
Determining A, B, or N

Is the salt NaF Acidic, basic or neutral in water?



- so NaF is **basic**

Is the salt NH_4NO_3 acidic, basic or neutral in aqueous solution?



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RJ Alexander



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Table of Contents Chem 12 Notes On Acids Bases Sss Chemistry

1. Understanding the eBook Chem 12 Notes On Acids Bases Sss Chemistry
 - The Rise of Digital Reading Chem 12 Notes On Acids Bases Sss Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 12 Notes On Acids Bases Sss Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 12 Notes On Acids Bases Sss Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 12 Notes On Acids Bases Sss Chemistry
 - Personalized Recommendations
 - Chem 12 Notes On Acids Bases Sss Chemistry User Reviews and Ratings
 - Chem 12 Notes On Acids Bases Sss Chemistry and Bestseller Lists
5. Accessing Chem 12 Notes On Acids Bases Sss Chemistry Free and Paid eBooks
 - Chem 12 Notes On Acids Bases Sss Chemistry Public Domain eBooks
 - Chem 12 Notes On Acids Bases Sss Chemistry eBook Subscription Services
 - Chem 12 Notes On Acids Bases Sss Chemistry Budget-Friendly Options
6. Navigating Chem 12 Notes On Acids Bases Sss Chemistry eBook Formats

- ePub, PDF, MOBI, and More
 - Chem 12 Notes On Acids Bases Sss Chemistry Compatibility with Devices
 - Chem 12 Notes On Acids Bases Sss Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem 12 Notes On Acids Bases Sss Chemistry
 - Highlighting and Note-Taking Chem 12 Notes On Acids Bases Sss Chemistry
 - Interactive Elements Chem 12 Notes On Acids Bases Sss Chemistry
 8. Staying Engaged with Chem 12 Notes On Acids Bases Sss Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 12 Notes On Acids Bases Sss Chemistry
 9. Balancing eBooks and Physical Books Chem 12 Notes On Acids Bases Sss Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 12 Notes On Acids Bases Sss Chemistry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chem 12 Notes On Acids Bases Sss Chemistry
 - Setting Reading Goals Chem 12 Notes On Acids Bases Sss Chemistry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chem 12 Notes On Acids Bases Sss Chemistry
 - Fact-Checking eBook Content of Chem 12 Notes On Acids Bases Sss Chemistry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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