

Acid-Base Titrations Pre-Lab Questions (10 Points)

Name _____ Lab Day _____

Lab Instructor _____ Date _____

Show calculations in space provided or on a separate sheet of paper.

1. In an experiment, 39.26 mL of 0.1062 M NaOH solution was required to titrate 17.54 mL of unknown acetic acid solution to a phenolphthalein end point. Calculate the molarity of the acetic acid solution, and the percent (by weight) of acetic acid in the solution (assuming its density to be 1.00 g/mL).

$$\begin{aligned}
 (M_2)(V_2) &= (M_1)(V_1) \\
 M_2 (17.54 \text{ mL}) &= (0.1062 \text{ M})(39.26 \text{ mL}) \\
 M_2 &= \frac{4.1654 \text{ M}^2\text{L}}{17.54 \text{ mL}} = 0.1110 \dots
 \end{aligned}$$

_____ 0.1111 _____ M _____ 11.11 _____ %

2. What is the proper disposal of all chemicals in this experiment?
 Discard all solutions, HCl solution, rinsing solutions, and titration solutions into your waste beaker. If your waste beaker gets filled, discard all solutions in the waste-acid base container.
3. List some the hazards associated with sodium hydroxide.
 Sodium hydroxide is toxic, corrosive, and can easily be absorbed by the skin.

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S Ashworth



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