

## 1.2 Mole Conversions Solutions

**PRACTICE EXAMPLES-** To complete the practice examples (1) write the type of question you have, (2) write the equation with units, (3) show the cancellation of units and (4) calculate the solution.

1. Determine how many atoms are present in 2.5 moles of Silicon? ( $15.05 \times 10^{23}$  atoms)

**MOLES → ATOMS** 
$$\frac{2.5 \text{ moles}}{1} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 15.05 \times 10^{23} \text{ atoms}$$

2. How many moles of Silicon is  $3.01 \times 10^{23}$  atoms? (.5 moles)

**ATOMS → MOLES** 
$$\frac{3.011 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ mole}}{6.022 \times 10^{23} \text{ atoms}} = .5 \text{ moles}$$

3. What is the mass of 4 moles of Iron? (224 grams)

**MOLES → GRAMS** 
$$\frac{4 \text{ moles}}{1} \times \frac{55.9 \text{ grams}}{\text{mole}} = 224 \text{ grams of C}$$

4. How many atoms are in 48.6 grams of Magnesium? ( $12.04 \times 10^{23}$  atoms)

**GRAMS → ATOMS** 
$$\frac{48.6 \text{ grams}}{1} \times \frac{1 \text{ mole}}{24.3 \text{ grams}} \times \frac{6.022 \times 10^{23} \text{ atoms}}{1 \text{ mole}} = 12.04 \times 10^{23} \text{ atoms}$$

5. How many grams are in  $18.06 \times 10^{23}$  atoms of Copper? (190.5 grams)

**ATOMS → GRAMS** 
$$\frac{18.06 \times 10^{23} \text{ atoms}}{1} \times \frac{1 \text{ mole}}{6.022 \times 10^{23} \text{ atoms}} \times \frac{63.5 \text{ grams}}{1 \text{ mole}} = 190.5 \text{ grams}$$

6. How many moles of Nickel is 176 grams? (3 moles)

**GRAMS → MOLES** 
$$\frac{176 \text{ grams of C}}{1} \times \frac{\text{mole}}{58.7 \text{ grams}} = 3 \text{ moles}$$

# Chemistry Conversion Practice Problems With Answers

**Margaret G. Kimble**



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## **Table of Contents Chemistry Conversion Practice Problems With Answers**

1. Understanding the eBook Chemistry Conversion Practice Problems With Answers
  - The Rise of Digital Reading Chemistry Conversion Practice Problems With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Conversion Practice Problems With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Conversion Practice Problems With Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Conversion Practice Problems With Answers
  - Personalized Recommendations
  - Chemistry Conversion Practice Problems With Answers User Reviews and Ratings
  - Chemistry Conversion Practice Problems With Answers and Bestseller Lists
5. Accessing Chemistry Conversion Practice Problems With Answers Free and Paid eBooks
  - Chemistry Conversion Practice Problems With Answers Public Domain eBooks
  - Chemistry Conversion Practice Problems With Answers eBook Subscription Services

- Chemistry Conversion Practice Problems With Answers Budget-Friendly Options
- 6. Navigating Chemistry Conversion Practice Problems With Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Conversion Practice Problems With Answers Compatibility with Devices
  - Chemistry Conversion Practice Problems With Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Conversion Practice Problems With Answers
  - Highlighting and Note-Taking Chemistry Conversion Practice Problems With Answers
  - Interactive Elements Chemistry Conversion Practice Problems With Answers
- 8. Staying Engaged with Chemistry Conversion Practice Problems With Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Conversion Practice Problems With Answers
- 9. Balancing eBooks and Physical Books Chemistry Conversion Practice Problems With Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Conversion Practice Problems With Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Conversion Practice Problems With Answers
  - Setting Reading Goals Chemistry Conversion Practice Problems With Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Conversion Practice Problems With Answers
  - Fact-Checking eBook Content of Chemistry Conversion Practice Problems With Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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