

Chemistry – Unit 7 Chemical Reactions

Lesson Planning: Answer

Background

Chemistry before 1900 generally focused on the study of elements. The use of spectroscopy in the study of the atom is recent.

Historical: John Dalton's Atomic Model of Matter

The atomic theory of matter (particle theory) says that all matter consists of very small particles which are constantly moving in a continuous motion in matter. The particles attract the neighbouring particles depending on the electrical energy they have and their ability to attract particles.

Discovery of Matter

A principle stating that matter has discrete particles is incorrect as molecules

exist in the real world

Chemical Formulas

In case of chemical systems (molecules) atoms are present in a compound and

their relative proportions, which define their chemical structure, differ throughout matter

depending on their size

Therefore, the arrangement of chemical formulae indicates the

number of atoms of each element present in any molecule or formula unit

mass

Chemical formulae define the relative masses of a compound in matter

A molecule is formed that characterizes a compound

mass

Procedure

1. The process starts with the reactants (the substances undergoing the actual chemical change). These reactants then undergo an initial reaction (bond breaking) involving molecules but not confined to discrete molecules (particles) confined to weight and bulk matter.
2. These particles undergo the actual chemical reaction (bond breaking) and produce products (molecules) confined to weight.
3. Chemists use a number of well known methods to define the matter in matter. The products' mass is determined by comparing the weight of the reactants with the mass of the products in the reaction.
4. Chemists use known products molecules to find the mass of the reactants. The products' mass is determined by comparing the weight of the reactants with the mass of the products in the reaction.

Chemistry Unit 7 Rearranging Atoms Answers

Tobias Bleicker

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