

Chapter Test II

A. Matching

Match each item in Column I with the correct description in Column II. Write the letter of the correct description in the blank.

Column I

- _____ 1. The volume of a fixed amount of gas varies inversely with the pressure of the gas at constant temperature.
- _____ 2. The volume of a fixed amount of gas varies directly with the absolute temperature.
- _____ 3. The volume of a fixed amount of gas varies directly with the absolute temperature at constant pressure.
- _____ 4. The volume of a fixed amount of gas varies directly with the absolute temperature at constant pressure and constant volume.
- _____ 5. A gas exerts pressure on the walls of its container because of the collisions of the gas molecules.
- _____ 6. The average kinetic energy of gas molecules is proportional to the absolute temperature.
- _____ 7. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant volume.
- _____ 8. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant pressure.
- _____ 9. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant volume.
- _____ 10. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant pressure.
- _____ 11. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant volume.
- _____ 12. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant pressure.
- _____ 13. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant volume.
- _____ 14. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant pressure.
- _____ 15. The average kinetic energy of gas molecules is proportional to the absolute temperature at constant volume.

Column II

- a. Boyle's law
- b. Charles' law
- c. Avogadro's law
- d. Gay-Lussac's law
- e. Dalton's law
- f. Graham's law
- g. Boyle's law
- h. Charles' law
- i. Dalton's law
- j. Graham's law
- k. Boyle's law
- l. Charles' law
- m. Dalton's law
- n. Graham's law
- o. Boyle's law
- p. Charles' law
- q. Dalton's law
- r. Graham's law

B. Multiple Choice

Circle the correct answer and write its letter in the blank.

- _____ 16. According to the kinetic theory of gases, the pressure of a gas is due to the collisions of the gas molecules with the walls of the container.
 - a. The molecules are in constant motion.
 - b. The molecules are in constant motion and they collide with the walls of the container.
 - c. The molecules are in constant motion and they collide with each other.
 - d. The molecules are in constant motion and they collide with the walls of the container.

Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers

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