

# Energy and Chemical Change

## Section 16.1 Energy

In your textbook, read about the nature of energy.

In the space at the left, write *true* if the statement is true; if the statement is false, change the italicized word or phrase to make it true.

- \_\_\_\_\_ 1. *Energy* is the ability to do work or produce heat.
- \_\_\_\_\_ 2. The law of conservation of energy states that energy *can be* created and destroyed.
- \_\_\_\_\_ 3. Chemical potential energy is energy stored in a substance because of its *composition*.
- \_\_\_\_\_ 4. *Heat* is a form of energy that flows from a warmer object to a cooler object.
- \_\_\_\_\_ 5. A calorie is the amount of energy required to raise the temperature of *one gram* of pure water by one degree Celsius.
- \_\_\_\_\_ 6. A calorie is the SI unit of heat and energy.
- \_\_\_\_\_ 7. The *specific heat* of a substance is the amount of heat required to raise the temperature of *one gram* of that substance by one degree Celsius.
- \_\_\_\_\_ 8. *Kinetic energy* is energy of motion.
- \_\_\_\_\_ 9. Chemicals participating in a chemical reaction contain *only potential energy*.
- \_\_\_\_\_ 10. One nutritional Calorie is equal to *100 calories*.
- \_\_\_\_\_ 11. One calorie equals *4,184 joules*.
- \_\_\_\_\_ 12. When a fuel is burned, some of its *chemical potential energy* is lost as heat.
- \_\_\_\_\_ 13. To convert kilojoules to joules, *divide* the number of kilojoules by 1000 joules/1 kilojoule.

Answer the following question. Show all your work.

14. If the temperature of a 500.0-g sample of liquid water is raised 2.00°C, how much heat is absorbed by the water? The specific heat of liquid water is 4.184 J/(g·°C).

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**D Keegan**



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