

# Answers to exam-style questions

## Chapter 3

- 1 C
- 2 B
- 3 A
- 4 D
- 5 A
- 6 C
- 7 A
- 8 B
- 9 D
- 10 B
- 11 B
- 12 D

13 a Pairs of electrons in the outer shell of the central atom repel each other and take up positions in space to be as far away from each other as possible in order to minimise this repulsion. Pairs of electrons may be bonding pairs and lone pairs; lone pairs cause greater repulsion than bonding pairs of electrons.

b i The Lewis structure for  $\text{PCl}_3$  is:



There are four pairs of electrons around the central atom. The basic shape (distribution of electron pairs) is tetrahedral, and because one of the pairs of electrons is a lone pair, the actual shape of the molecule is trigonal pyramidal. The bond angle will be slightly less than that in a tetrahedron ( $109.5^\circ$ ), so predictions in the range  $100\text{--}108^\circ$  will generally be accepted.

ii The Lewis structure for  $\text{CO}_2$  is:



The four electrons in each double bond are constrained to be in the same region of space, so there are two negative charge centres around the C. The shape is linear and the bond angle is  $180^\circ$ .

c  $\text{CO}_2$  is a linear molecule and so, although each individual  $\text{C}=\text{O}$  bond is polar, the dipoles cancel out.



Because there is a lone pair of electrons on the S in  $\text{SO}_2$ , the molecule is bent and the dipoles do not cancel out – one end of the molecule is  $\delta^+$  and the other is  $\delta^-$ .



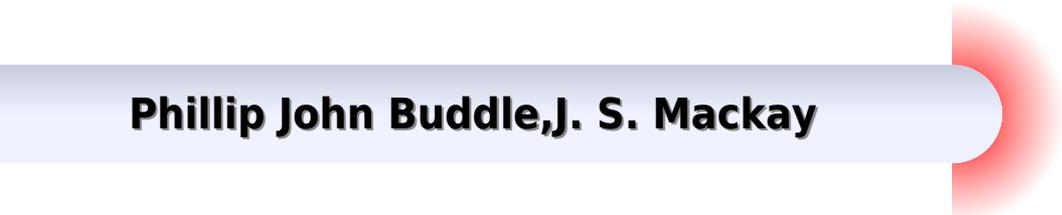
d i  $\text{C}\equiv\text{O}$ :

CO has a triple bond between the C and the O, whereas  $\text{CO}_2$  has double bonds between the C and O atoms; the C–O bond in CO is shorter than that in  $\text{CO}_2$ ; there are six electrons between the atoms in a triple bond but only four electrons between the atoms in a double bond; with more electrons between the atoms, there is greater attraction between the nuclei and the bonding electrons.

- 14 a Sodium oxide has a giant ionic structure in which the ions are held in a lattice structure by strong electrostatic forces of attraction between positive and negative ions. As the electrostatic attractions between the oppositely charged ions are strong, a lot of energy is required to break them, and the melting point is high.  $\text{Na}_2\text{O}$  does not conduct electricity when solid, as the ions are held tightly in the lattice structure. When molten, the  $\text{Na}^+$  and  $\text{O}^{2-}$  ions are free to move, and so the liquid conducts electricity.
- b Sodium and magnesium have metallic bonding, which is the electrostatic attraction between positively charged ions and the delocalised electrons.  $\text{Mg}^{2+}$  has a higher positive charge than  $\text{Na}^+$ , so there is a stronger attraction for the delocalised electrons.  $\text{Mg}^{2+}$  is smaller than  $\text{Na}^+$ , so the nucleus of the ion is closer to the delocalised electrons and the attraction is stronger. In Mg there are more delocalised electrons per atom, and therefore more attractions holding the lattice together.
- c These are all covalent molecular substances. The strengths of van der Waals' forces increase with relative molecular mass.  $\text{AsH}_3$  has a higher relative

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**Phillip John Buddle, J. S. Mackay**



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