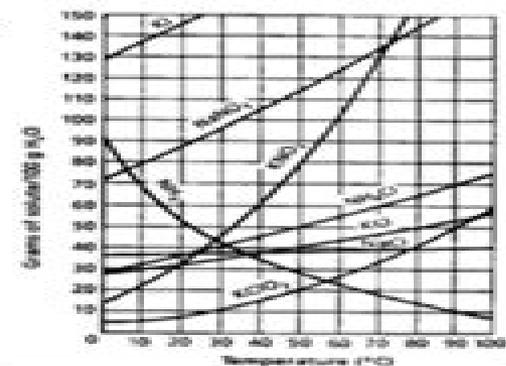


Name _____
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SOLUTIONS WORKSHEET

A. Solubility curves – refer to the graph to answer the following questions.



1. What is the solubility of NH₄Cl at 70°C? **60 g**
2. Calculate the mass of NH₄Cl that would dissolve in 200 g of water at 90°C.

70 g @ 100g water → 140 g

3. At 0°C, 50.0 g NH₄Cl are added to 100 g water. To what temperature does the solution have to be raised in order to achieve a saturated solution of NH₄Cl_(aq)?

50°C

4. Is NH₃ a (s), (l), or (g)? Explain.

GAS – it's solubility decreases with increase in temperature. Opposite for solids

5. Determine how much water is required to dissolve 35.0 g NH₄Cl at 70°C in order to make a saturated solution. Show your work.

$$\frac{60}{100} = \frac{35}{x} \quad x = 58.3 \text{ g H}_2\text{O}$$

6. If a saturated solution of NH₃ at 15°C is warmed to 40°C, how much ammonia will "fall out" of the solution?

@ 15°C – 60 g soluble @ 40°C – 35 g soluble 25 g will fall out

B. MOLARITY

1. What is molarity a measure of? What is the formula for molarity? What are the units of molarity?

Concentration. $M = n / V$ mol/L or M

2. Calculate the molarity of a 1.5 L solution that has 0.45 mol solute dissolved.

0.30 - M

3. Calculate the grams of sodium sulfate that are required to prepare 1.50 L of 0.25 M solution.

53.3 g Na₂SO₄

4. Calculate the volume of solution required to make 0.024 M solution using 15.0 g of calcium chloride.

V = 5.6 L

5. Describe in detail the steps you would take to prepare 1.50 L of 0.5 M solution of sodium chloride.

n = 0.75 mol → 43.5 g NaCl

- 1) measure out 43.5 g NaCl
- 2) put NaCl into graduated cylinder
- 3) add enough water to make final volume 1.50 L

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Puja Mehta



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