



CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90	91	92	93	94	95	96	97	98	99
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PERIODIC TABLE

- GROUPS**
 - Elements in a vertical column are grouped together. Substances of a particular group share similar chemical and physical properties.
 - There are 18 vertical columns in the periodic table.
 - There is a break in the number of groups between groups 8 and 9.
- PERIODS**
 - Elements in a horizontal row are grouped together.
 - There are 7 horizontal rows in the periodic table.
 - There is a break in the number of periods between periods 6 and 7.
- THE BLOCKS OF THE PERIODIC TABLE**
 - The periodic table is divided into four main blocks: s-block, p-block, d-block, and f-block.

ATOMIC STRUCTURE AND THE PERIODIC TABLE

- ATOMIC STRUCTURE**
 - Atoms are made up of three subatomic particles: electrons, protons, and neutrons.
 - Electrons are negatively charged, protons are positively charged, and neutrons are neutral.
 - The number of protons in an atom is equal to the number of electrons, making the atom electrically neutral.
- THE PERIODIC TABLE**
 - The periodic table is arranged in order of increasing atomic number.
 - Elements with similar properties are grouped together.

$1 - \frac{1}{n^2} = R_H \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$

where $R_H = 1.097 \times 10^7 \text{ m}^{-1}$, n_1 and n_2 are integers representing the initial and final energy levels of the electron in the transition.

IONIZATION ENERGY, ELECTRONEGATIVITY AND ELECTROPOSITIVE CHARACTER



Ionization energy is the energy required to remove an electron from an atom. It increases across a period and decreases down a group.

ELECTRONEGATIVITY OF ATOMS

- The tendency of an atom to attract a shared pair of electrons towards itself.
- Increases across a period and decreases down a group.
- Fluorine is the most electronegative element.
- Francium is the least electronegative element.



THE WAVE MECHANICAL MODEL

- Electrons are not in fixed orbits but exist in regions called orbitals.
- Orbitals are regions of space where the probability of finding an electron is high.
- The energy of an electron in an orbital depends on the principal quantum number (n).

- The energy of an electron in an orbital is given by $E_n = -\frac{13.6}{n^2} \text{ eV}$.
- The energy difference between two orbitals is given by $\Delta E = E_2 - E_1$.
- The wavelength of light emitted or absorbed during a transition is given by $\lambda = \frac{hc}{\Delta E}$.
- The frequency of light is given by $\nu = \frac{c}{\lambda}$.

MOLECULES

- Molecules are formed by the combination of atoms.
- The number of atoms in a molecule is given by the molecular formula.
- The relative molecular mass of a molecule is given by $M_r = \sum (\text{atoms} \times \text{atomic mass})$.
- The number of moles of a substance is given by $n = \frac{m}{M_r}$.

PERIODIC TABLE

- The periodic table is arranged in order of increasing atomic number.
- Elements with similar properties are grouped together.
- The periodic table is divided into four main blocks: s-block, p-block, d-block, and f-block.
- The periodic table is used to predict the properties of elements.
- The periodic table is used to identify elements.
- The periodic table is used to determine the number of protons and neutrons in an atom.
- The periodic table is used to determine the number of electrons in an atom.
- The periodic table is used to determine the number of valence electrons in an atom.
- The periodic table is used to determine the number of shells in an atom.
- The periodic table is used to determine the number of subshells in an atom.
- The periodic table is used to determine the number of orbitals in an atom.
- The periodic table is used to determine the number of nodes in an orbital.
- The periodic table is used to determine the number of nodes in a molecule.
- The periodic table is used to determine the number of nodes in a crystal lattice.
- The periodic table is used to determine the number of nodes in a polymer chain.
- The periodic table is used to determine the number of nodes in a protein chain.
- The periodic table is used to determine the number of nodes in a DNA molecule.
- The periodic table is used to determine the number of nodes in a virus particle.
- The periodic table is used to determine the number of nodes in a cell membrane.
- The periodic table is used to determine the number of nodes in a cell wall.
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MOLECULES

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- The relative molecular mass of a molecule is given by $M_r = \sum (\text{atoms} \times \text{atomic mass})$.
- The number of moles of a substance is given by $n = \frac{m}{M_r}$.
- The number of molecules in a sample is given by $N = n \times N_A$.
- The number of atoms in a sample is given by $N_{\text{atoms}} = n \times N_A \times \text{atoms per molecule}$.
- The number of ions in a sample is given by $N_{\text{ions}} = n \times N_A \times \text{ions per molecule}$.
- The number of electrons in a sample is given by $N_{\text{electrons}} = n \times N_A \times \text{electrons per molecule}$.
- The number of protons in a sample is given by $N_{\text{protons}} = n \times N_A \times \text{protons per molecule}$.
- The number of neutrons in a sample is given by $N_{\text{neutrons}} = n \times N_A \times \text{neutrons per molecule}$.
- The number of shells in a sample is given by $N_{\text{shells}} = n \times N_A \times \text{shells per atom}$.
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- The number of orbitals in a sample is given by $N_{\text{orbitals}} = n \times N_A \times \text{orbitals per atom}$.
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JR Anderson



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