

Worksheet #1: Significant Figures & Dimensional Analysis

Directions: For each problem below, write the equation and show your work. Be sure to **box your final answer**.

Part 1: Solve the following problems using scientific notation and rounding to the appropriate value.

a. $300,234,800 \rightarrow 3.0 \times 10^8$

b. $450,500 \rightarrow 4.5 \times 10^5$

c. $0.000027100 \rightarrow 2.71 \times 10^{-5}$

d. $-0.000000 \rightarrow -3.5 \times 10^{-7}$

\times / \div
 → Sig figs in the result is the same as the in the least precise measurement.

Part 2: Solve the following problems, and show your final answer with the appropriate number of significant figures.

a. $1.24056 + 75.80 \rightarrow 77.04$

b. $10^3 + 99(34.5 - 20) \rightarrow 1.21$

c. $0.8897 \times 2.15 + 0.0001 \rightarrow 1.9$

d. $43.0 \times 8.0 + 89.32(75) \rightarrow 411.2$

e. $(2.88 + 53) \div (23,000 - 0.77) \rightarrow 9.9 \times 10^{-7}$

$+ / -$
 → same # of decimal places as the least precise measurement.

Part 3: For each problem below, show your work. Always use units and box in your final answer.

a. The density of pure silver is 10.5 g/cm^3 at 20°C . If 5.25 g of pure silver pellets are added to a graduated cylinder containing 11.2 mL of water, to what volume level will the water in the cylinder rise?

$d = 10.5 \text{ g/cm}^3$
 $m = 5.25 \text{ g}$
 $v = ?$



$d = \frac{m}{v}$
 $10.5 = \frac{5.25}{x}$
 $x = 0.5 \text{ cm}^3$

$11.2 \text{ mL} + 0.5 \text{ mL}$
 $= 11.7 \text{ mL}$

b. The density of air at ordinary atmospheric pressure and 25°C is 1.19 g/L . What is the mass, in kilograms, of the air in a room that measures $12.5 \times 15.5 \times 8.0 \text{ ft}$?

$d = 1.19 \text{ g/L}$

$1550 \text{ ft}^3 \cdot \left(\frac{28.32 \text{ L}}{1 \text{ ft}^3} \right) = 43896 \text{ L}$

52.2 kg

28.32 L
 (28.32 is used for more info)

$v = 12.5 \times 15.5 \times 8.0$
 $= 1550 \text{ ft}^3$

$d = \frac{m}{v} \rightarrow 1.19 = \frac{x}{43896} \Rightarrow 52236.24 \text{ g}$

c. An aluminum block has a density of 2.70 g/mL . If the mass of the block is 24.60 g , find the volume of the substance.

$d = 2.70 \text{ g/mL}$

$m = 24.60 \text{ g}$

$v = ?$

$d = \frac{m}{v} \rightarrow 2.70 = \frac{24.60}{x} = 9.111 \approx 9.11 \text{ mL}$

Part 4: Convert the following measurements to the desired unit.

a. $0.0001 \text{ cm} = 0.50 \text{ mm}$

c. $1 \text{ ft} = 12 \text{ in}$

b. $1872 \text{ mg} = 0.001872 \text{ kg}$

d. $3.6 \times 10.3 \text{ ks} = 370 \text{ s}$

$\frac{1 \text{ min}}{60 \text{ sec}}$
 $\frac{1 \text{ hr}}{60 \text{ min}}$
 $\frac{1 \text{ day}}{24 \text{ hr}}$

Chemistry States Of Matter Packet Answers Key

Tao Wei



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Table of Contents Chemistry States Of Matter Packet Answers Key

1. Understanding the eBook Chemistry States Of Matter Packet Answers Key
 - The Rise of Digital Reading Chemistry States Of Matter Packet Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry States Of Matter Packet Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry States Of Matter Packet Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry States Of Matter Packet Answers Key
 - Personalized Recommendations
 - Chemistry States Of Matter Packet Answers Key User Reviews and Ratings
 - Chemistry States Of Matter Packet Answers Key and Bestseller Lists

5. Accessing Chemistry States Of Matter Packet Answers Key Free and Paid eBooks
 - Chemistry States Of Matter Packet Answers Key Public Domain eBooks
 - Chemistry States Of Matter Packet Answers Key eBook Subscription Services
 - Chemistry States Of Matter Packet Answers Key Budget-Friendly Options
6. Navigating Chemistry States Of Matter Packet Answers Key eBook Formats
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 - Chemistry States Of Matter Packet Answers Key Compatibility with Devices
 - Chemistry States Of Matter Packet Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry States Of Matter Packet Answers Key
 - Highlighting and Note-Taking Chemistry States Of Matter Packet Answers Key
 - Interactive Elements Chemistry States Of Matter Packet Answers Key
8. Staying Engaged with Chemistry States Of Matter Packet Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry States Of Matter Packet Answers Key
9. Balancing eBooks and Physical Books Chemistry States Of Matter Packet Answers Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry States Of Matter Packet Answers Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry States Of Matter Packet Answers Key
 - Setting Reading Goals Chemistry States Of Matter Packet Answers Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry States Of Matter Packet Answers Key
 - Fact-Checking eBook Content of Chemistry States Of Matter Packet Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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