

AP Biology Midterm Review Study Guide



- **Midterm exam date:** period 1: January 22, 2016, 8:00am – 10:00am
period 3: January 25, 2016, 8:00am – 10:00am
- **Format:** 40 Multiple Choice Questions, math questions, and free response questions
- This midterm will count as your midterm exam grade and your final exam grade. These 2 grades will be averaged with your 5th marking period project grade to determine your 5th marking period grade.
- Use this review sheet to help guide you during your preparation for this exam. **NOTE:** This list is not all-inclusive.

2 - Biochemistry

- What are the life processes?
- What determines the different types of bonds (Nonpolar Covalent, Polar Covalent, Ionic, Hydrogen)?
- Mass # , atomic #, isotopes, ions, electron configuration
- Protons (mass, +1, nucleus), neutron (mass, +0, nucleus), electron (no mass, -1, orbitals, valence e- involved in bonds)

2 – Water

- What types of bonds do you find within a water molecule? * * * between water molecules?
- What atom in water is very electronegative?
- Properties: excellent solvent, high heat capacity—moderating influence on temperature, evaporative cooling, ice floats - insulation, strong cohesion & surface tension, strong adhesion
- PH scale 0-14 acid = 0-(7) high [H⁺], low [OH⁻] neutral = 7 base (7)-14, low [H⁺], high [OH⁻]
- What are buffers? Why are they necessary?

3 – Carbon / Organic Molecules

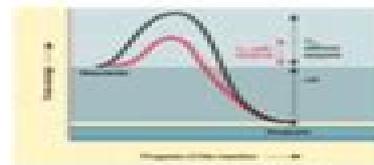
- What are some properties of the carbon (organic) atom?
- Functional Groups: structural formulas, properties, and examples of molecules they are found in:
 - Amino, central carbonyl, terminal carbonyl, carboxyl, hydroxyl, methyl, phosphate, sulfhydryl
- Isomers = same number of atoms, different properties

3- Macromolecules

- What is the difference between dehydration synthesis and hydrolysis?
 - $C_6H_{12}O_6 + C_6H_{12}O_6 = C_{12}H_{22}O_{11} \text{ (disaccharide)} + H_2O$
- Know macromolecule monomers vs. polymers, bonds, and examples:
 - Carbohydrates: monosaccharides, disaccharides, polysaccharides, sugars, starch, glycogen, cellulose, chitin, glycosidic linkage bonds
 - Proteins - structure, transport, defense, enzymes, amino acids, dipeptides, polypeptides, proteins, peptide bonds, 1^o, 2^o, 3^o, 4^o levels of structure
 - Lipids - energy storage, structure, hormones, groups: triglycerides (fats, saturated C-C, unsaturated/kinky C=C), phospholipids, steroids, (cholesterol, sex hormones), ester linkage bonds.
 - Nucleic acids- genetic information storage, structure: nucleotides (pentose sugar, nitrogenous base (A, T, C, G, U), negatively charged phosphate groups, DNA, RNA, hydrogen bonds between bases, covalent bonds in backbone)

6 - Enzymes/ Metabolism

- What do the first and second laws of thermodynamics state?
- What is the structure and purpose of ATP?
- What happens when high-energy phosphate bonds are broken in ATP?
- $\Delta G_{\text{reaction}}$ (Change in Energy of Reaction) = G_{products} (Energy of Products) – $G_{\text{reactants}}$ (Energy of Reactants)
- Define: Catabolism, Hydrolysis, Exergonic, Negative Delta G
- What macromolecule are enzymes made of? What do they do to a substrate? Can they be reused? What do enzymes do to the activation energy of a reaction?
- Function: metabolic catalysts, lock & key model, induced fit model, "-ase", substrate specific
- Factors affecting enzyme action: pH, temperature, coenzymes, cofactors, substrate concentration, enzyme concentration
- Activators (ex: allosteric) vs. inhibitors (ex: competitive or non-competitive)
- Allosteric site vs. active site
- Negative feedback loop / feedback inhibition
- Be able to label & interpret a Rates of Reaction Diagram ->



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Deborah T. Goldberg



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