

## AP Biology Summer Assignment Chapter 2 Quiz 2016-17

### Multiple Choice

Identify the choice that best completes the statement or answers the question.

- Which six elements provide most of the mass of the human body?
  - C, O, N, H, P, S
  - K, O, N, H, P, C
  - C, O, Na, H, P, S
  - C, O, N, H, F, Mg
- Which particles determine how an atom interacts with other atoms?
  - Protons
  - Neutrons
  - Electrons
  - Nuclei
- Refer to the figure below.



A neutral atom of nitrogen is shown. When reacting with another atom to achieve a stable configuration of electrons, this atom will likely

- gain one electron in its outer shell to make its total number of electrons equal to eight.
  - gain three electrons in its outer shell to make the number of electrons in this shell equal to eight.
  - lose the five electrons in its outer shell, which leaves a full inner shell of two electrons.
  - lose one electron in its outer shell, which leaves an even number of six electrons in the outer shell.
  - not react because the atom is already stable with an equal number of protons and electrons.
- A covalent bond
    - may be polar, but cannot be nonpolar.
    - may contain electrons shared unequally between two nuclei.
    - forms when two atoms share a single electron.
    - links an atom that donates electrons to an atom that gains the donated electrons.
  - The electronegativity of an atom is a relative measure of the
    - number of electrons in an atom.
    - number of electrons in the atom's outermost electron shell.
    - difference between the number of atoms and the number of protons.
    - affinity an atom has for electrons and its ability to capture additional electrons.
  - Refer to the figure below.  
 $\text{H}_3\text{C}-\text{CH}_3$   
What type of chemical bond connects the two carbon atoms in the molecule shown?
    - Nonpolar covalent bond
    - Asymmetric bond
    - Polar covalent bond
    - Hydrogen bond
  - Refer to the figure below.  
 $\text{H}_3\text{C}-\text{OH}$   
What type of chemical bond connects the carbon and oxygen atoms in the molecule shown?
    - Nonpolar covalent bond
    - Asymmetric bond
    - Polar covalent bond
    - Hydrogen bond
  - Which of the following is responsible for water's high heat capacity, high surface tension, strong cohesion, and strong adhesion properties?
    - The size of the water molecule
    - Hydrogen bonding between molecules
    - Water's low mass
    - Water's abundance on Earth
    - Water's ability to change state

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**Y Pai**



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