

## AP Chemistry Test (Chapter 13)

### Multiple Choice (30%)

- 1) Which one best describes the  $K_c$  for this reaction?



- A)  $K_c = \frac{[A]^3[D]^2}{[E]^4[D]^2}$       B)  $K_c = \frac{[B]^3[E]^4[G]^2}{[A]^3[D]^2[C]^5}$   
C)  $K_c = \frac{[E]^4[B]^3}{[A]^3[D]^2}$       D)  $K_c = \frac{[A]^3[D]^2[C]^5}{[B]^3[E]^4[G]^2}$   
E) None of these is correct

- 2) Please consider the gas phase reaction:



Which one would decrease the yield of  $H_2$  ?

- A) High T & high P      B) Low T & low P  
C) High T & low P      D) Low T & high P  
3)  $AlBr_3$  dissolves in water, according to this mechanism:



What effect would the addition of  $Na_2PO_4$  have on this equilibrium?

- A) Q becomes larger & the reaction creates more products.  
B) Q becomes larger & the reaction creates more reactants.  
C) Q becomes smaller & the reaction creates more products.  
D) Q becomes smaller & the reaction creates more reactants.  
E) There is no net effect on the equilibrium position.  
4) Based upon the following information, which one is true about the reaction mixture?



$$P (NH_3) = 2.19 \text{ atm} \quad P (N_2) = 4.33 \text{ atm} \quad P (H_2) = 0.72 \text{ atm}$$

- A) The product concentrations are too high & the reaction will shift left to establish equilibrium.  
B) The product concentrations are too high & the reaction will shift right to establish equilibrium.  
C) The reactant concentrations are too high & the reaction will shift right to establish equilibrium.  
D) The reactant concentrations are too high & the reaction will shift left to establish equilibrium.

# Ap Chemistry Chapter 13 Test

**K Payea**



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