

## Key Ideas

- ▶ What is the difference between protons, neutrons, and electrons?
- ▶ What do atoms of an element have in common with other atoms of the same element?
- ▶ Why do isotopes of the same element have different atomic masses?
- ▶ What unit is used to express atomic mass?

## Key Terms

proton  
neutron  
atomic number  
mass number  
isotope  
unified atomic mass unit  
mole

## Why It Matters

Radioisotopes emit energy when they decay. To diagnose and treat diseases, doctors use this property of radioisotopes to track where in the body certain atoms go.

Less than 100 years after Dalton published his atomic theory, scientists determined that atoms consisted of smaller particles, such as the electron. In this section, you will learn more about the particles inside the atom.

## What Is in an Atom?

Atoms are made up of various subatomic particles. To understand the chemistry of most substances, however, we need to study only three of these particles. ▶ **The three main subatomic particles are distinguished by mass, charge, and location in the atom.** Figure 1 compares these particles.

At the center of each atom is a small, dense nucleus. The nucleus is made of **protons**, which have a positive charge, and **neutrons**, which have no charge. Protons and neutrons are almost identical in size and mass. Moving around outside the nucleus is a cloud of very tiny, negatively charged electrons. The mass of an electron is much smaller than that of a proton or neutron.

Figure 1 Subatomic Particles

Particle	Charge	Mass (kg)	Location in the atom
Proton	+1	$1.67 \times 10^{-27}$	in the nucleus
Neutron	0	$1.67 \times 10^{-27}$	in the nucleus
Electron	-1	$9.11 \times 10^{-31}$	outside the nucleus

**proton** (PROH) is a subatomic particle that has a positive charge and that is located in the nucleus of an atom.

**neutron** (NEO) is a subatomic particle that has no charge and that is located in the nucleus of an atom.

## Focus

In this section, students compare the three basic subatomic particles (protons, electrons, neutrons). They learn how atomic number, mass number, and molar mass are used to represent properties of atoms.

## Bellringer

Use the Bellringer transparency to prepare students for this section.

## Demonstrate

## Counting Large Numbers by Mass

Put students into groups. Each group will need a large container, identical items to fill the container (such as beans), and a balance. Ask students to find the following: the mass of the empty container ( $m_{\text{cont}}$ ), the mass of a few ( $x$ ) of the items ( $m_{x \text{ items}}$ ), and the mass of the container with all of the items in it ( $m_{\text{tot}}$ ). Tell each group to use their data to calculate the number of small items without counting them. (First, find the mass of one item by dividing the mass of a given number of items by that number. Then find the mass of all of the items by subtracting the mass of the container from the total mass. Finally, the number of unknown items ( $n$ ) equals the mass of all of the items divided by the mass of one item.) Tell students that in this section they will see how this method can be used to “count atoms.” **LE Kinesthetic**

## Key Resources

 **Teaching Transparencies**  
C8 Nucleus  
TM12 Subatomic Particles  
C9 Isotopes  
TM13 Mole-Mass Conversion

 **Visual Concepts**  
Atomic Number  
Mass Number  
Isotopes and Nuclides  
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The Mole

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Modeling Isotopes

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# Chapter 4 Section 2 Notes Riusechchemistry Weebly

**Katrin Zwirglmaier**



**Chapter 4 Section 2 Notes Riusehchemistry Weebly:**

## **Chapter 4 Section 2 Notes Riusechchemistry Weebly** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the ability of words has become more evident than ever. They have the ability to inspire, provoke, and ignite change. Such could be the essence of the book **Chapter 4 Section 2 Notes Riusechchemistry Weebly**, a literary masterpiece that delves deep to the significance of words and their impact on our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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