

**General Chemistry I**  
**Practice Test with Answer Key**  
**Chemical Reaction**

**Test I: True or False.** Write **True** if the statement is correct and write **False** if the statement is wrong.

- \_\_\_\_\_ 1) chemical reactions can be easily analyzed if they are written through chemical equations  
\_\_\_\_\_ 2) Reactants are substances produced after a chemical reaction.  
\_\_\_\_\_ 3) Chemical Equation is the shorthand for describing the course of chemical reaction  
\_\_\_\_\_ 4) Chemical Reaction is a process in which one or more substances are changed into new substance with the same physical and chemical properties  
\_\_\_\_\_ 5) (g) is an indication that the substance is in gas form  
\_\_\_\_\_ 6) Main key to balance an equation is to write the correct subscript before the chemical symbol  
\_\_\_\_\_ 7) Converting the word equation into a chemical equation is one of the pre-step in balancing  
\_\_\_\_\_ 8)  means heat is applied during the reaction  
\_\_\_\_\_ 9) (aq) means aqueous, substance is dissolved in ammonia  
\_\_\_\_\_ 10) The water we drink every day, the sugar found in bread, the carbon dioxide gas produced from burning a piece of paper are examples produced as the result of chemical reactions.

**Test II: Multiple Choices.** Choose the **best answer** among the choices. Write the letter of the correct answer before the number.

- \_\_\_\_\_ 1) Substances involved before a chemical reaction  
a. product b. reactant c. compound d. elements  
\_\_\_\_\_ 2) The end or final materials in a chemical reaction is known as  
a. Reactant b. product c. substance c. quotient  
\_\_\_\_\_ 3) What is the meaning of + used in a chemical equation?  
a. yields b. reacts with c. heat is applied d. gas  
\_\_\_\_\_ 4) In the reaction  $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$ , oxygen gas is  
a. subscript b. product c. reactant d. coefficient  
\_\_\_\_\_ 5) Which equation is correctly balanced?  
a.  $\text{C} + \text{O}_2 \rightarrow 2 \text{CO}$  b.  $\text{C} + \text{O}_2 \rightarrow \text{CO}$  c.  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  d.  $2 \text{C} + \text{O}_2 \rightarrow \text{CO}$   
\_\_\_\_\_ 6) In the equation  $2 \text{C}_2\text{H}_2(\text{g}) + 5 \text{O}_2 \rightarrow 4 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ , the coefficient for carbon dioxide is  
a. 2 b. 5 c. (g) d. 4  
\_\_\_\_\_ 7) All of the following are chemical reactions EXCEPT

- \_\_\_\_\_ a.  $\text{H}_2\text{O}(\text{l}) \rightarrow \text{H}_2\text{O}(\text{g})$   
\_\_\_\_\_ b.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ c.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ d.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ e.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ f.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ g.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ h.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ i.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ j.  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$

- \_\_\_\_\_ 8)  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ 9)  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ 10)  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ 11)  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$   
\_\_\_\_\_ 12)  $\text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NaCl}(\text{s})$

# Chemical Reaction Test And Answer

**Jo Allan Beran**



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