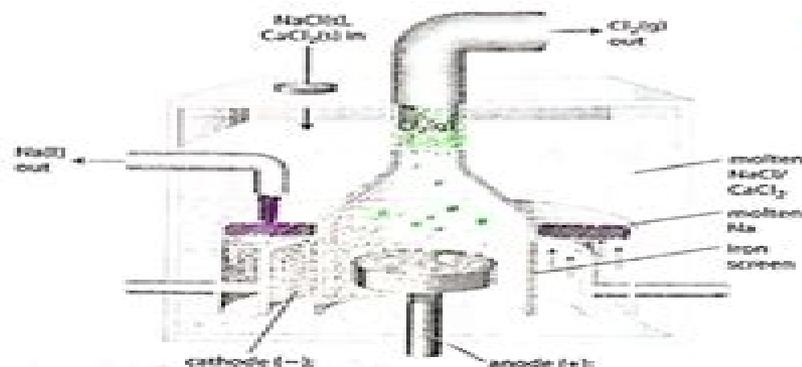


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Electrolysis of sodium chloride (in the industry)

Industrial electrolysis of molten Sodium Chloride requires extreme temperatures.

- Calcium Chloride is added to the mixture to lower the melting point of Sodium Chloride.
- Molten sodium metal floats in the mixture because of its low density.
- The anode is carbon
- The cathode is iron
- Chlorine gas is collected separately



3.

4. electrolysis of Aluminum oxide (molten)

Electrodes: graphite (carbon)

Electrolyte: molten Alumina (Al₂O₃)

Alumina = Purified Bauxite

Which substance is added to the molten alumina? cryolite

Ions Present: Al³⁺, O²⁻

Half equations:

In the Cathode (Al³⁺ + 3e⁻ → Al) × 4

In the Anode (2O²⁻ → O₂ + 4e⁻) × 3

Overall Equation 4 Al³⁺ + 6 O²⁻ → 4 Al + 3 O₂

Observations:

Cathode: The Molten Aluminum can be observed

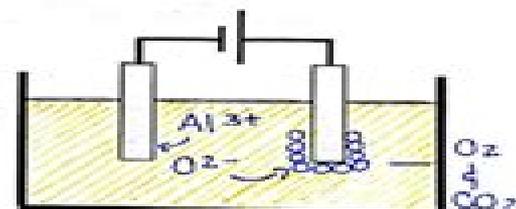
Anode: Bubbles of Oxygen gas will appear

Why is it necessary to change the electrodes periodically?

Because the anodes will burn out. Anode reacts with O₂ to form CO₂

Which is the purpose of the cryolite?

To lower the melting point of the Aluminum Oxide



Chemistry Electrochemistry Study Guide Answers Key

Rachel S Tattersall



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Table of Contents Chemistry Electrochemistry Study Guide Answers Key

1. Understanding the eBook Chemistry Electrochemistry Study Guide Answers Key
 - The Rise of Digital Reading Chemistry Electrochemistry Study Guide Answers Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Electrochemistry Study Guide Answers Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Electrochemistry Study Guide Answers Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Electrochemistry Study Guide Answers Key
 - Personalized Recommendations
 - Chemistry Electrochemistry Study Guide Answers Key User Reviews and Ratings
 - Chemistry Electrochemistry Study Guide Answers Key and Bestseller Lists

5. Accessing Chemistry Electrochemistry Study Guide Answers Key Free and Paid eBooks
 - Chemistry Electrochemistry Study Guide Answers Key Public Domain eBooks
 - Chemistry Electrochemistry Study Guide Answers Key eBook Subscription Services
 - Chemistry Electrochemistry Study Guide Answers Key Budget-Friendly Options
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 - Chemistry Electrochemistry Study Guide Answers Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Electrochemistry Study Guide Answers Key
 - Highlighting and Note-Taking Chemistry Electrochemistry Study Guide Answers Key
 - Interactive Elements Chemistry Electrochemistry Study Guide Answers Key
8. Staying Engaged with Chemistry Electrochemistry Study Guide Answers Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Electrochemistry Study Guide Answers Key
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Electrochemistry Study Guide Answers Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Chemistry Electrochemistry Study Guide Answers Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Electrochemistry Study Guide Answers Key
 - Fact-Checking eBook Content of Chemistry Electrochemistry Study Guide Answers Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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