

Stoichiometry Worksheet # 1

1) Glucose is used as a source of energy by the human body. The overall reaction in the body is



Calculate the number of moles of oxygen needed to oxidize 12.5 g of $\text{C}_6\text{H}_{12}\text{O}_6$.

2) Ammonia is synthesized from hydrogen and nitrogen according to the following equation.



If an excess of nitrogen is reacted with 3.41 moles of hydrogen gas, how many grams of ammonia (NH_3) can be produced?

3) Assume that in the decomposition of potassium chlorate, KClO_3 ,



80.5 moles of O_2 form. How many grams of potassium chlorite, the other product would be formed?

4) In a single displacement reaction,



9.23 moles of aluminum react with excess hydrochloric acid. How many moles of hydrogen will be produced?

5) The compound "Cisplatin" $\text{PtCl}_2(\text{NH}_3)_2$, has been found to be effective in treating some types of cancer. It can be synthesized using the following reaction



A. How many moles of "cisplatin" can be produced from 2.50 g K_2PtCl_4 ?

B. How many moles of NH_3 would be needed if 3.8 moles of K_2PtCl_4 were present?

Chapter 12 Stoichiometry D Reading Answers

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Table of Contents Chapter 12 Stoichiometry D Reading Answers

1. Understanding the eBook Chapter 12 Stoichiometry D Reading Answers
 - The Rise of Digital Reading Chapter 12 Stoichiometry D Reading Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 12 Stoichiometry D Reading Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 12 Stoichiometry D Reading Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 12 Stoichiometry D Reading Answers
 - Personalized Recommendations

- Chapter 12 Stoichiometry D Reading Answers User Reviews and Ratings
- Chapter 12 Stoichiometry D Reading Answers and Bestseller Lists
- 5. Accessing Chapter 12 Stoichiometry D Reading Answers Free and Paid eBooks
 - Chapter 12 Stoichiometry D Reading Answers Public Domain eBooks
 - Chapter 12 Stoichiometry D Reading Answers eBook Subscription Services
 - Chapter 12 Stoichiometry D Reading Answers Budget-Friendly Options
- 6. Navigating Chapter 12 Stoichiometry D Reading Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 12 Stoichiometry D Reading Answers Compatibility with Devices
 - Chapter 12 Stoichiometry D Reading Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 12 Stoichiometry D Reading Answers
 - Highlighting and Note-Taking Chapter 12 Stoichiometry D Reading Answers
 - Interactive Elements Chapter 12 Stoichiometry D Reading Answers
- 8. Staying Engaged with Chapter 12 Stoichiometry D Reading Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 12 Stoichiometry D Reading Answers
- 9. Balancing eBooks and Physical Books Chapter 12 Stoichiometry D Reading Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 12 Stoichiometry D Reading Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 12 Stoichiometry D Reading Answers
 - Setting Reading Goals Chapter 12 Stoichiometry D Reading Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 12 Stoichiometry D Reading Answers
 - Fact-Checking eBook Content of Chapter 12 Stoichiometry D Reading Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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