

# CHAPTER 14

## CHEMICAL EQUILIBRIUM

14.13  $K_c = \frac{[B]}{[A]}$

- (1) With  $K_c = 10$ , products are favored at equilibrium. Because the coefficients for both A and B are one, we expect the concentration of B to be 10 times that of A at equilibrium. Choice (a) is the best choice with 10 B molecules and 1 A molecule.
- (2) With  $K_c = 0.10$ , reactants are favored at equilibrium. Because the coefficients for both A and B are one, we expect the concentration of A to be 10 times that of B at equilibrium. Choice (d) is the best choice with 10 A molecules and 1 B molecule.

You can calculate  $K_c$  in each case without knowing the volume of the container because the mole ratio between A and B is the same. Volume will cancel from the  $K_c$  expression. Only moles of each component are needed to calculate  $K_c$ .

14.14 Note that we are comparing similar reactions at equilibrium – two reactants producing one product, all with coefficients of one in the balanced equation.

- (a) The reaction,  $A + C \rightleftharpoons AC$  has the largest equilibrium constant. Of the three diagrams, there is the most product present at equilibrium.
- (b) The reaction,  $A + D \rightleftharpoons AD$  has the smallest equilibrium constant. Of the three diagrams, there is the least amount of product present at equilibrium.

14.15 When the equation for a reversible reaction is written in the opposite direction, the equilibrium constant becomes the reciprocal of the original equilibrium constant.

$$K' = \frac{1}{K} = \frac{1}{4.17 \times 10^{-34}} = 2.40 \times 10^{33}$$

14.16 The problem states that the system is at equilibrium, so we simply substitute the equilibrium concentrations into the equilibrium constant expression to calculate  $K_c$ .

**Step 1:** Calculate the concentrations of the components in units of mol/L. The molarities can be calculated by simply dividing the number of moles by the volume of the flask.

$$[H_2] = \frac{2.50 \text{ mol}}{12.0 \text{ L}} = 0.208 \text{ M}$$

$$[S_2] = \frac{1.35 \times 10^{-5} \text{ mol}}{12.0 \text{ L}} = 1.13 \times 10^{-6} \text{ M}$$

$$[H_2S] = \frac{8.70 \text{ mol}}{12.0 \text{ L}} = 0.725 \text{ M}$$

# Chapter 14 Chemical Equilibrium Quiz

**Conrad L. Stanitski**



## Chapter 14 Chemical Equilibrium Quiz:

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first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

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