

## CHAPTER 16 REVIEW

### *Reaction Energy*

#### SECTION 1

**SHORT ANSWER** Answer the following questions in the space provided.

1. For elements in their standard state, the value of  $\Delta H_f^\circ$  is \_\_\_\_\_.
2. The formation and decomposition of water can be represented by the following thermochemical equations:



- \_\_\_\_\_ a. Is energy being taken in or is it being released as liquid  $\text{H}_2\text{O}$  decomposes?
- \_\_\_\_\_ b. What is the appropriate sign for the enthalpy change in this decomposition reaction?

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

3. \_\_\_\_\_ If 200. g of water at  $20^\circ\text{C}$  absorbs 41 840 J of energy, what will its final temperature be?
  
4. \_\_\_\_\_ Aluminum has a specific heat of  $0.900 \text{ J}(\text{g}\cdot^\circ\text{C})$ . How much energy in kJ is needed to raise the temperature of a 625 g block of aluminum from  $30.7^\circ\text{C}$  to  $82.1^\circ\text{C}$ ?
  
5. The products in a reaction have an enthalpy of  $-458 \text{ kJ/mol}$ , and the reactants have an enthalpy of  $658 \text{ kJ/mol}$ .
  - \_\_\_\_\_ a. What is the value of  $\Delta H$  for this reaction?

# Chapter 16 Review Reaction Energy Section 1 Short Answer

**Terry C. Jones**



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