

10. States of Matter : Gaseous and liquid states

10.1 Introduction:

We have learnt that substances exist in one of the three main states of matter. The three distinct physical forms of a substance are Solid, Liquid, and Gas.



Can you recall?

Water exists in the three different forms solid ice, liquid water and gaseous vapours.

Key points of differentiation between the three states can be understood as given in Table 10.1.



Fig. 10.1 : Different States of Water

Three states of matter are interconvertible by exchange of Heat as given below:



Table 10.1: Distinguishing points between Solid, Liquid and Gas

Sr. No.	Points	Solid	Liquid	Gas
1	Microscopic view Mean atomic/ molecular separation	 Mean separation = $3-5A^{\circ}$	 Mean separation = $3-10A^{\circ}$	 Mean separation $> 5A^{\circ}$
2	Arrangement of particles (atoms/ molecules)	Particles are tightly held, and have regular arrangement of atoms/ molecules	Particles are loosely packed, irregular arrangement of particles	Particles are more loosely packed, highly irregular arrangement
3	Movement of particles	Particles cannot move freely as they occupy fixed positions.	Particles can move a small distance within the liquid	Particles are in continuous random motion.
4	Shape and volume	Has definite shape and volume	Takes the shape of the container and has definite volume.	Takes the shape and the volume of its container.
5	Intermolecular space	Very small Intermolecular space	Moderate Intermolecular space	Large Intermolecular space.
6	Effect of a small change in temperature	Volume change is small	Moderate effect on volume change	Volume change significantly high.
7	Compression or Expansion	Practically Non-compressible	Small Compressibility	Compressible

Chapter 10 States Of Matter Test

Rachel Sandford



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