

Chapter 16 Solubility and Complex Ion Equilibria

Section 16.1: Solubility Equilibria and the Solubility Product

◆ 1. What is Solubility Equilibrium?

- When an ionic solid (like AgCl) is put in water, some dissolves into ions.
- At the same time, ions in the water can recombine and form the solid again.
- Eventually, it reaches **equilibrium**: the rate of dissolving = rate of crystallizing.

Example: $AgCl_s \rightleftharpoons Ag^+_{aq} + Cl^-_{aq}$

◆ 2. Solubility Product (K_{sp})

- **K_{sp}** is the equilibrium constant for this dissolving.
- For AgCl: $K_{sp} = [Ag^+][Cl^-]$

👉 Notice: The solid (AgCl) is left out because its concentration is constant.

◆ 3. Solubility vs. K_{sp}

- **Solubility** = how many moles per liter dissolve.
- **K_{sp}** = a number from the ion concentrations at equilibrium.

⚠ Important: A larger K_{sp} doesn't always mean higher solubility because the ion ratios matter.

Example:

- AgCl: $K_{sp} = 1.6 \times 10^{-10}$, solubility $\approx 1.3 \times 10^{-5}$.
 - Ag₃PO₄: $K_{sp} = 1.8 \times 10^{-18}$, solubility $\approx 1.6 \times 10^{-5}$.
Even though Ag₃PO₄ has a smaller K_{sp}, its solubility is close because of the 3:1 ratio of ions.
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Chapter 16 Solubility And Complex Ion Equilibria

Zhe Xu



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Table of Contents Chapter 16 Solubility And Complex Ion Equilibria

1. Understanding the eBook Chapter 16 Solubility And Complex Ion Equilibria
 - The Rise of Digital Reading Chapter 16 Solubility And Complex Ion Equilibria
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 16 Solubility And Complex Ion Equilibria
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 16 Solubility And Complex Ion Equilibria
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 16 Solubility And Complex Ion Equilibria
 - Personalized Recommendations
 - Chapter 16 Solubility And Complex Ion Equilibria User Reviews and Ratings
 - Chapter 16 Solubility And Complex Ion Equilibria and Bestseller Lists
5. Accessing Chapter 16 Solubility And Complex Ion Equilibria Free and Paid eBooks
 - Chapter 16 Solubility And Complex Ion Equilibria Public Domain eBooks
 - Chapter 16 Solubility And Complex Ion Equilibria eBook Subscription Services

- Chapter 16 Solubility And Complex Ion Equilibria Budget-Friendly Options
- 6. Navigating Chapter 16 Solubility And Complex Ion Equilibria eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 16 Solubility And Complex Ion Equilibria Compatibility with Devices
 - Chapter 16 Solubility And Complex Ion Equilibria Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 16 Solubility And Complex Ion Equilibria
 - Highlighting and Note-Taking Chapter 16 Solubility And Complex Ion Equilibria
 - Interactive Elements Chapter 16 Solubility And Complex Ion Equilibria
- 8. Staying Engaged with Chapter 16 Solubility And Complex Ion Equilibria
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 16 Solubility And Complex Ion Equilibria
- 9. Balancing eBooks and Physical Books Chapter 16 Solubility And Complex Ion Equilibria
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 16 Solubility And Complex Ion Equilibria
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 16 Solubility And Complex Ion Equilibria
 - Setting Reading Goals Chapter 16 Solubility And Complex Ion Equilibria
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 16 Solubility And Complex Ion Equilibria
 - Fact-Checking eBook Content of Chapter 16 Solubility And Complex Ion Equilibria
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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