

Worksheet 4D: Solution Calculations

Supplemental Instruction
Iowa State University

Leader: Sara

Course: Chem177

Instructor: Bonaccorsi

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Dilution

1) In lab you mix 23.5 grams of copper (II) Nitrate (molar mass = 187.6) into 193 mL of water.

a. What is the molarity of the solution?

$$\frac{23.5 \text{ g Cu(NO}_3)_2 \cdot (1 \text{ mol}/187.6 \text{ g})}{.193 \text{ L H}_2\text{O}} = .649 \text{ M}$$

b. You then dilute it to .500 L, at which point the concentration of the solution is what?

$$M_2 = .649 \cdot .193 / .500 = .251 \text{ M}$$

Limiting Reagent/ Solution Stoichiometry

2) $3\text{Au(NO}_3)_3 \text{ (aq)} + 2\text{Li}_3\text{PO}_4 \text{ (aq)} \rightarrow 6\text{LiNO}_3 \text{ (aq)} + \text{Au}_3(\text{PO}_4)_2 \text{ (s)}$

a. What is the net ionic equation for the reaction?



b. If you start with 10.0 mL of 6.00 M Gold (II) Nitrate and you add 113 mL of water what is the new concentration?

$$10.0 \text{ mL} \cdot 6.00 \text{ mol/L} = (113 \text{ mL} + 10.0) \cdot M_2 \\ M_2 = .488 \text{ M}$$

c. You react it with .0500 moles of Lithium phosphate. How many moles of solid is produced?

$$(.488 \text{ mol gold/L} \cdot .123 \text{ L}) \cdot (1 \text{ mole solid}/3 \text{ moles gold}) = .0200 \text{ mol solid} \\ .0500 \text{ mol Li} \cdot (1 \text{ moles solid}/2 \text{ moles Li}) = .0250 \text{ moles solid} \\ \text{Gold limiting}$$

3) In lab you are performing a titration to find the concentration of a 25.6 mL solution of H_2SO_4 . You start by dissolving 8.56 g solid KOH (molar mass = 56.1 g/mol) in 50.0 mL water. You then use 11.7 mL of that solution to titrate the H_2SO_4 (aq).

a. Write out the molecular and net ionic reaction?



b. What is the concentration of the basic solution you made?

$$\frac{8.56 \text{ g KOH} \cdot (1 \text{ mol KOH} / 56.1 \text{ g KOH})}{.0500 \text{ L H}_2\text{O}} = 3.05 \text{ M}$$

c. What is the unknown concentration of the acid solution?

$$\frac{.0117 \text{ L base} \cdot 3.05 \text{ moles base} \cdot 1 \text{ mole acid}}{1 \text{ L base} \cdot 2 \text{ moles base} \cdot .0256 \text{ L base}} = .697 \text{ M}$$

Stoichiometry Phet Lab Answers

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