

Chapter Test B

Chapter Chemical Equations and Reactions

Write the full name (including units) for the term at the end of the sentence. Please show your work. Please write answers without answers with quantities.

- The molar volume of an ideal gas at standard temperature and pressure is 22.7 L mol^{-1} .
 - gas
 - pressure
 - standard temperature
 - volume
- The volume of a gas at standard temperature is 1.0 L and the pressure is 101.3 kPa .
 - volume
 - pressure
 - standard temperature
 - number of particles
- An unknown gas has a molar mass of 44.01 g mol^{-1} .
 - gas
 - density in g L^{-1}
 - number
 - volume
- The molar mass of a gas is 44.01 g mol^{-1} . The gas is at standard temperature and pressure.
 - standard temperature and pressure
 - density in g L^{-1}
 - number of particles
 - number of particles
- The molar mass of a gas is 44.01 g mol^{-1} . The gas is at standard temperature and pressure.
 - density in g L^{-1}
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 - number of particles
 - number of particles

Modern Chemistry Chapter 7 Test B Answers

Marcel A. Müller



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