

## Module Overview

This is an introductory course for students who are interested in

### Objectives

The student will be able to:

- 1. Calculate the equilibrium constant for a reaction from the standard free energy of formation of the reactants and products.
- 2. Calculate the standard free energy of formation of a reactant or product from the equilibrium constant for a reaction and the standard free energy of formation of the other reactants or products.
- 3. Calculate the standard free energy of formation of a reactant or product from the equilibrium constant for a reaction and the standard free energy of formation of the other reactants or products.

### Module 1 - Chemical Equilibrium

$$aA + bB \rightleftharpoons cC + dD$$

1. Calculate the equilibrium constant  $K_c$  for the reaction:
  - a.  $2H_2(g) + O_2(g) \rightleftharpoons 2H_2O(g)$   $K_c = 4.0$
  - b.  $2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$   $K_c = 0.25$
2. Calculate the equilibrium constant  $K_c$  for the reaction:
  - a.  $2H_2(g) + O_2(g) \rightleftharpoons 2H_2O(g)$   $K_c = 4.0$
  - b.  $2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$   $K_c = 0.25$
3. Calculate the equilibrium constant  $K_c$  for the reaction:
  - a.  $2H_2(g) + O_2(g) \rightleftharpoons 2H_2O(g)$   $K_c = 4.0$
  - b.  $2H_2O(g) \rightleftharpoons 2H_2(g) + O_2(g)$   $K_c = 0.25$

	$K_c$ (molar)	$K_p$ (atmos)	$K_c$ (partial)	$K_p$ (atmos)
For a reaction involving gases, the equilibrium constant $K_c$ is calculated by dividing the product of the concentrations of the products by the product of the concentrations of the reactants.	1	1	1	1
For a reaction involving gases, the equilibrium constant $K_p$ is calculated by dividing the product of the partial pressures of the products by the product of the partial pressures of the reactants.	1	1	1	1
For a reaction involving gases, the equilibrium constant $K_c$ is calculated by dividing the product of the concentrations of the products by the product of the concentrations of the reactants.	1	1	1	1

# Mole Ratio Chemistry Lab Answer Key Thefl

**Yi-Tong Ma**



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