

Titration Calculations & Answers

1. Use the information to determine the concentration of the hydrochloric acid.
- A 25 cm³ sample of hydrochloric acid is sucked into a pipette and transferred into a 250 cm³ volumetric flask. The solution is made up to the mark.
 - 25 cm³ of the diluted acid is transferred into a conical flask using a pipette.
 - A burette is used to neutralise the acid with 0.1 mol dm⁻³ sodium hydroxide.

Hydrochloric acid reacts with sodium hydroxide according to the equation:



- a. The average titre of the sodium hydroxide solution was 30.00 cm³. Calculate the number of moles in the average titre.

$$n = c \times (v/1000) = 0.100 \times (30.00/1000) = 0.003 \text{ mol (1)}$$

0.003 mol (1)

- b. Determine the number of moles in the diluted sample of hydrochloric acid, and hence the concentration of the diluted acid.

$$\text{Moles in diluted sample} = 0.003 \text{ mol (1)}$$

$$c = n/(v/1000) = 0.003/(25/1000) = 0.12 \text{ mol dm}^{-3} \text{ (1)}$$

0.12 mol dm⁻³ (2)

- c. Calculate the concentration of the undiluted hydrochloric acid in mol dm⁻³.

$$\text{Dilution} = 250/25 = 10$$

$$10 \times 0.003 = 0.12 \text{ mol dm}^{-3} = 1.2 \text{ mol dm}^{-3} \text{ (1)}$$

1.2 mol dm⁻³ (1)

- d. Calculate the concentration of the hydrochloric acid in g dm⁻³.

$$\text{Mr} = 1 + 35.5 = 36.5 \text{ (1)}$$

$$m = n \times \text{Mr} = 1.2 \times 36.5 = 43.8 \text{ g dm}^{-3} \text{ (1)}$$

43.8 g dm⁻³ (2)

Titration Questions And Answers

D Keegan



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