

Vol. of 1 mol of ideal gas at STP/NTP =  $22.4 \text{ L}$

उत्स नान एरिया  
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoga ans me.

\* Significant fig  
• 0 to the left of the 1st non 0 digit are sig. not sig.  
eg  $0.0012 \Rightarrow 2$   
• Terminal 0 are not sig. if there is no decimal pt.  
eg  $1000 \Rightarrow 1$  fig.  
• 0 at end is sig. if there is decimal pt.  
• Exponential term are not considered sig.

अंश - 3/131  
 $a \times b = c$  } - add/subst  
is least amount  
→ 3 same sabse kam fig hoga vo sig. fig ans me hoga

\* LAW OF MASS CONSV  
• In a react<sup>n</sup> → actual mass of reactant = mass of prod  
Rakme zitna use hua.

\* Except<sup>n</sup> → Nuclear Rxn (E = mc<sup>2</sup>)  
• PppWed<sup>n</sup> → Balanced Rxn

# MOLE CONCEPT

\* LAW OF CONST PROPORT<sup>n</sup>  
• Elements in a comp are in fixed ratio by mass.  
eg  $\text{CO}_2$   
 $12 \quad 32 \text{g} \Rightarrow 3:8$   
Not valid for isotope.  
% of C =  $\frac{3}{11} \times 100$

Q. cal. the mass of C req to form 132g of  $\text{CO}_2$ .  
 $\Rightarrow \text{C} + 2\text{O} \rightarrow \text{CO}_2$   
 $12 \quad 32 \quad 44 \text{g}$   
44g  $\text{CO}_2$  ke lie 12g C  
∴ 132g ke lie  
 $= \frac{12}{44} \times 132 = 36 \text{g}$

\* Law of Multiple Prop<sup>n</sup>  
• elements should be same or same fixed mass  
eg.  $\text{CO}$  &  $\text{CO}_2$  1:2 ratio  
Other eg.  $\text{N}_2\text{O}$ ,  $\text{N}_2\text{O}_3$ ,  $\text{N}_2\text{O}_5$   
except<sup>n</sup> →  $\text{N}_2\text{O}_3$  &  $\text{N}_2\text{O}_5$

\* Average molar mass  
• Same (P, T, R, V) of molecules of gases some (mass eg! not of molar)  
 $\text{Xe}_2, \text{Xe}_3, \text{Xe}_4$   
 $2 : 3 : 4$  molecules  
 $2 : 3 : 4$  atoms  
[different]

\* Gay Lussac's law  
CONDINS: +T, P some  
• at least 2 gases.  
 $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$   
 $1 : 1 : 2$   
• Vol. R, L

\* DALTON'S ATOMIC THEORY  
• Atoms → Neither created nor destroyed

\* MOLE CONCEPT  
Atomic No. (P) → mass no. (P+N)  
• Mass of Proton =  $1.67 \times 10^{-27} \text{ kg}$   
• Neutron =  $1.67 \times 10^{-27} \text{ kg}$   
• Electron =  $9.1 \times 10^{-31} \text{ kg}$   
• MP = 1837  $1/e$

\* Atomic mass unit (amu)  
 $1 \text{ amu} = \frac{1}{12} \times \text{Mass of } ^{12}\text{C}$   
 $1 \text{ amu} = 1.67 \times 10^{-27} \text{ kg}$   
eg Mass of  $^{12}\text{C} = 12 \text{ amu}$   
• Relative atomic mass  
• Unitless  
• eg. mass. A.M. of C = 12

\* Molecular Mass  
•  $\text{H}_2\text{SO}_4 = 98 \text{ amu}$   
↳ Mass of 1 molecule of  $\text{H}_2\text{SO}_4$   
• FORMULA MASS: for ionic comp.

\* MOLE  
Amount of substance.  
• 1 mole = No particles  
( $6.022 \times 10^{23}$ )  
 $1 \text{ amu} \leq 1.67 \times 10^{-27} \text{ kg}$   
 $\text{Na} = 6.022 \times 10^{23}$   
 $1 \text{ amu} \times \text{Na} = 1 \text{ g}$

# Mole Concept Theory Notes Pdfslibforme

**Wolfgang Guggemos**



**Mole Concept Theory Notes Pdfslibforme:**

## **Mole Concept Theory Notes Pdfslibforme** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Mole Concept Theory Notes Pdfslibforme**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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