

## 15.4 DETERMINING ENTHALPIES OF REACTION OR FORMATION [ΔH°]

A first method used to determine enthalpies of reaction is the standard enthalpy of formation.

Others are calorimetry and Hess's law.

For example, the formation equation and the standard enthalpy of formation of carbon dioxide are:  
 $C_{(s)} + O_{2(g)} \rightarrow CO_{2(g)}$     ΔH° = -393.5 kJ/mol

Steps for writing Formation Equations:

1. Write one mole of product in the state that has been specified
2. Write the reactant elements in their standard states
3. Choose reaction coefficients for the reactants to give a balanced equation yielding one mole of product.

Example 1:

Write the equation for formation of liquid ethanol

1. Start with one mole of product:



2. Write the reactant elements in their standard states:



3. Adjust the equation to yield one mole of product.



Example 2:

What is the formation equation for liquid carbonic acid?



Solution:



Example 3:



ΔH°<sub>f</sub> is always for 1 mole of the product being formed

In general ΔH°<sub>r</sub> = (ΔH°<sub>f</sub> compounds formed) - (ΔH°<sub>f</sub> element<sub>1</sub> + ΔH°<sub>f</sub> element<sub>2</sub> + .....ΔH°<sub>f</sub>)

ENTHALPY OF REACTION: ΔH°<sub>r</sub> =

is equal to the difference between the sum of enthalpies of the formation of products and that of the reactants.

ΔH°<sub>r</sub> = enthalpy of product - enthalpy of reactants:

$$\Delta H^\circ_r = \sum \Delta H^\circ_f \text{ Products} - \sum \Delta H^\circ_f \text{ Reactants}$$

Ex:    aB + cD → eE + fF

$$\Delta H^\circ_r = [e\Delta H^\circ_f(E) + f\Delta H^\circ_f(F)] - [a\Delta H^\circ_f(A) + c\Delta H^\circ_f(C)]$$

Try the examples.

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**Raffaella Di Napoli**



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## **Table of Contents Nelson Chemistry 12 Solutions**

1. Understanding the eBook Nelson Chemistry 12 Solutions
  - The Rise of Digital Reading Nelson Chemistry 12 Solutions
  - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Chemistry 12 Solutions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Nelson Chemistry 12 Solutions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Chemistry 12 Solutions
  - Personalized Recommendations
  - Nelson Chemistry 12 Solutions User Reviews and Ratings
  - Nelson Chemistry 12 Solutions and Bestseller Lists
5. Accessing Nelson Chemistry 12 Solutions Free and Paid eBooks

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- Nelson Chemistry 12 Solutions Public Domain eBooks
  - Nelson Chemistry 12 Solutions eBook Subscription Services
  - Nelson Chemistry 12 Solutions Budget-Friendly Options
6. Navigating Nelson Chemistry 12 Solutions eBook Formats
    - ePub, PDF, MOBI, and More
    - Nelson Chemistry 12 Solutions Compatibility with Devices
    - Nelson Chemistry 12 Solutions Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Nelson Chemistry 12 Solutions
    - Highlighting and Note-Taking Nelson Chemistry 12 Solutions
    - Interactive Elements Nelson Chemistry 12 Solutions
  8. Staying Engaged with Nelson Chemistry 12 Solutions
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Nelson Chemistry 12 Solutions
  9. Balancing eBooks and Physical Books Nelson Chemistry 12 Solutions
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Nelson Chemistry 12 Solutions
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Nelson Chemistry 12 Solutions
    - Setting Reading Goals Nelson Chemistry 12 Solutions
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Nelson Chemistry 12 Solutions
    - Fact-Checking eBook Content of Nelson Chemistry 12 Solutions
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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