

## Double Replacement Reactions and Solubility

**Post-Lab Questions:** Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1)	$\text{Na}_2\text{CO}_3(\text{aq}) + \text{Al}(\text{NO}_3)_3(\text{aq}) \rightarrow 6\text{NaNO}_3(\text{aq}) + \text{Al}(\text{OH})_3(\text{s})$	None
2)	$\text{NaCl}(\text{aq}) + \text{NH}_4\text{NO}_3(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{NH}_4\text{Cl}(\text{aq})$	None
3)	$\text{NaI}(\text{aq}) + \text{AgNO}_3(\text{aq}) \rightarrow \text{NaNO}_3(\text{aq}) + \text{AgI}(\text{s})$	$\text{Ag}^+(\text{aq}) + \text{I}^-(\text{aq}) \rightarrow \text{AgI}(\text{s})$
4)	$\text{Na}_3\text{PO}_4(\text{aq}) + \text{Cu}(\text{NO}_3)_2(\text{aq}) \rightarrow 6\text{NaNO}_3(\text{aq}) + (\text{Cu}_3(\text{PO}_4)_2(\text{s}))$	$3\text{Cu}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Cu}_3(\text{PO}_4)_2(\text{s})$

Use your observations table to answer the following. **ALL ions must include a symbol AND a charge or it is wrong!**

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are \_\_\_\_\_

b) Salts of these anions may be described as SOLUBLE except for those with the cations \_\_\_\_\_

6a) The symbols and charges of the anion that produced the most (>5) precipitates are \_\_\_\_\_

b) Salts of these anions may be described as INSOLUBLE except for those with the cations \_\_\_\_\_

7. Although you did not test the bromide ion ( $\text{Br}^-$ ), would you expect most bromide salts ( $\text{AlBr}_3$ ,  $\text{BaBr}_2$ ,  $\text{CaBr}_2$ , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. \_\_\_\_\_

8. Explain your answer in #7 \_\_\_\_\_

9. What cation combined with bromide is an exception to the rule in #7? \_\_\_\_\_

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

# Double Replacement Reaction Lab Answers

**Camilla Rothe**



**Double Replacement Reaction Lab Answers:**

## Embracing the Song of Phrase: An Mental Symphony within **Double Replacement Reaction Lab Answers**

In some sort of eaten by displays and the ceaseless chatter of instant interaction, the melodic splendor and emotional symphony produced by the published term usually fade in to the backdrop, eclipsed by the persistent sound and disturbances that permeate our lives. Nevertheless, set within the pages of **Double Replacement Reaction Lab Answers** a charming fictional value full of natural emotions, lies an immersive symphony waiting to be embraced. Crafted by a masterful musician of language, this captivating masterpiece conducts readers on a psychological journey, well unraveling the hidden songs and profound impact resonating within each carefully crafted phrase. Within the depths of this touching analysis, we shall examine the book is main harmonies, analyze their enthralling writing style, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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