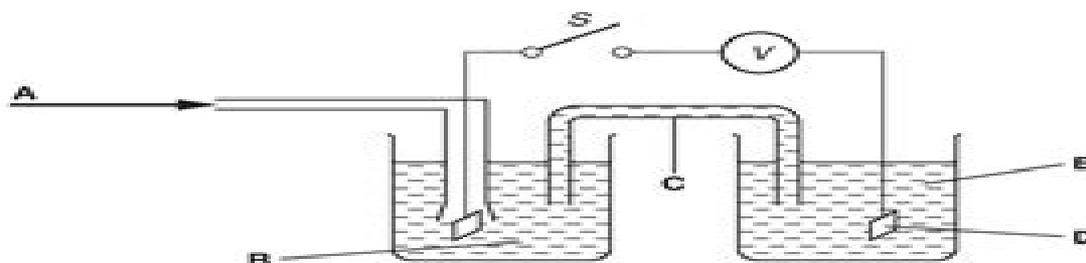


ELECTROCHEMISTRY WS 1

- 1 Chlorine gas and iron(II) ions react together in aqueous solution as follows.



- (a) The following diagram shows the apparatus needed to measure the E_{cell}° for the above reaction.



- (i) In the spaces below, identify what the five letters A – E in the above diagram represent.

A chlorine at $P_{\text{Cl}_2} = 1 \text{ atm}$ and 298 K
 B $1 \text{ mol dm}^{-3} \text{ NaCl}$
 C salt bridge
 D Platinum
 E $1 \text{ mol dm}^{-3} \text{ Fe}^{2+} \text{ aq}$ and $1 \text{ mol dm}^{-3} \text{ Fe}^{3+} \text{ aq}$

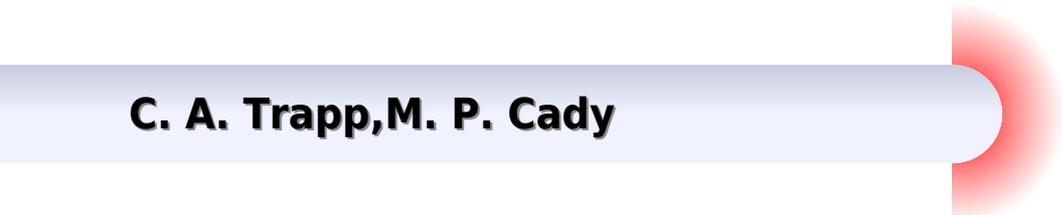
- (ii) Use the *Data Booklet* to calculate the E_{cell}° for this reaction, and hence decide which direction (left to right, or right to left) electrons would flow through the voltmeter V when switch S is closed.



[7]

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C. A. Trapp, M. P. Cady



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