



How To Start Ai Writing Assistant 2025 For Experts

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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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from one species to another the species that loses electrons is said to be oxidized while the species that gains electrons is said to be reduced

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web describe redox reactions in terms of electron transfer recall that oxidation involves loss of electrons and reduction involves gain of electrons represent simple redox reactions using half equations identify oxidising and reducing agents for simple redox reactions understand the reactivity of metals and metal displacement reaction

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you $23 \text{ Fe} + \text{CO} \rightarrow \text{CO}_2 + \text{Fe}$
oxidation $\text{CO} \rightarrow \text{CO}_2 + 2\text{e}^-$ reduction $\text{Fe} + 2\text{e}^- \rightarrow \text{Fe}$ $24 \text{ 3 Ag} + \text{Ni} \rightarrow \text{Ni}_3 + 3 \text{ Ag}$ oxidation reduction $25 \text{ Cu} + \text{Pb} \rightarrow \text{Pb}_2 + \text{Cu}$ oxidation reduction 26 O_2

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 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$ $5 \text{ C} + \text{O}_2 \rightarrow 5 \text{ C} + \text{O}_2$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ $\text{Sr} \rightarrow \text{Sr}^{2+}$ oxidized reducing agent $\text{O}_2 \rightarrow \text{O}^{2-}$ reduced ox ag $2\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$ $\text{Li} \rightarrow \text{Li}^+$ oxidized red ag $\text{S} \rightarrow \text{S}^{2-}$ reduced ox ag $3 \text{ 2Cs} + \text{Br}_2$

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