

Cracking the Periodic Table Code

Why aren't the elements listed in alphabetical order?

Why?

As chemists go, the periodic table is a bit odd. It's not square. Large portions of the table appear to be missing at the top. It's not organized alphabetically so elements can be found easily. But to a chemist, the periodic table is a very powerful tool. The periodic table is organized by properties, both chemical and physical. These properties relate to the electronic structure of the atoms of each element. In today's activity, you are going to study how the ground state electron configurations and structure of atoms are related to the shape and organization of the periodic table.

Model 1 – Blank Periodic Table

The diagram shows a blank periodic table grid. The rows are numbered 1 to 7 on the left. The columns are numbered 1 to 18 at the bottom. The grid is divided into four blocks:

- s-block:** The first two columns (columns 1 and 2).
- p-block:** The last six columns (columns 13, 14, 15, 16, 17, and 18).
- d-block:** The middle ten columns (columns 3 through 12).
- f-block:** A separate row of ten columns located below the main grid, labeled 'f' on the left.

- Obtain a card with electron configurations for your group as assigned by your teacher.
 - Record your team number _____ and team name _____.
 - For each element on your card, write the end of the configuration (the unsharpened portion).
- What is similar about the last orbital notation appearing at the end of the configuration for each element in your set?

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David Kirk



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