



Which statement correctly describes the energy changes that occur in the forward reaction?

- A) The activation energy is 10. kJ and the reaction is endothermic.
- B) The activation energy is 10. kJ and the reaction is exothermic.
- C) The activation energy is 50. kJ and the reaction is endothermic.
- D) The activation energy is 50. kJ and the reaction is exothermic.

9. Given the reaction at equilibrium:



The equilibrium point will shift to the right if the pressure is

- A) increased and the temperature is increased
- B) increased and the temperature is decreased

formation of 1 mole of  $H_2O(l)$

- A)  $-285.8 \text{ kJ}$
- B)  $+285.8 \text{ kJ}$
- C)  $-571.6 \text{ kJ}$
- D)  $+571.6 \text{ kJ}$

11. As the number of effective collisions between reacting particles increases, the rate of reaction

- A) decreases
- B) increases
- C) remains the same

12. Which statement must be true when solution equilibrium occurs?

- A) The solution is at STP.
- B) The solution is supersaturated.
- C) The concentration of the solution remains constant.
- D) The masses of the dissolved solute and the undissolved solute are equal.

13. Given the reaction at equilibrium:



The addition of a catalyst will

- A) shift the equilibrium to the right
- B) shift the equilibrium to the left
- C) increase the rate of forward and reverse reactions equally
- D) have no effect on the forward or reverse reactions

14. Which equation shows an increase in entropy?

- A)  $CO_2(g) \rightarrow CO_2(s)$
- B)  $CO_2(l) \rightarrow CO_2(g)$

# Practice Problems Solutions Kinetics And Equilibrium

**Stefan Bosworth**



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