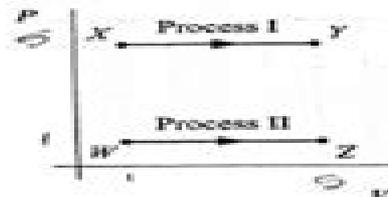




1. For each of the following parts, determine whether there exists an ideal gas process that satisfies the given conditions. If so, describe an experiment (e.g., one involving a cylinder sealed with a movable, frictionless piston) that is consistent with the given conditions. Give an example from tutorial if possible. If not, explain how you know that such a process does not exist.
  - a. There is heat transfer, but the temperature of the gas does not change (i.e.,  $Q \neq 0, \Delta T = 0$ ).  
 In the scenario via container of gas via piston, if heat is applied but the temperature doesn't change the piston would have to move up to allow more work to be done increasing the volume.
  - b. There is no heat transfer, but the temperature of the gas changes (i.e.,  $Q = 0, \Delta T \neq 0$ ).  
 A container with a gas & a piston, the piston would move down and compress the gas & increase the temperature.
  - c. There is no heat transfer, but work is done on the gas (i.e.,  $Q = 0, W \neq 0$ ).  
 The piston compressing the gas is work done on the gas.
  - d. There is heat transfer, but no work is done on the gas (i.e.,  $Q \neq 0, W = 0$ ).  
 The piston remains locked and an outside source heats the container.

2. One mole of an ideal gas is confined to a container with a movable piston. The questions below refer to the processes shown on the PV diagram at right. Process I is a change from state X to state Y at constant pressure. Process II is a change from state W to state Z at a different constant pressure.



- a. Rank states W-Z according to the temperature, from largest to smallest. If any temperatures are equal, state so explicitly. Explain.

$W(1X) = 1, X(1X) = 5, Y(5X) = 25, Z(5X) = 5,$

$$Y > X = Z > W$$

- b. For each process, state whether the piston moves inward, moves outward, or does not move. Explain.

$W \rightarrow Z$ : Piston moves outward (volume increases, pressure stays the same)  
 $X \rightarrow Y$ : Piston moves outward (volume increases, pressure remains the same)

- c. Based on your answer to part b, state whether the following quantities are positive, negative, or zero. Explain your reasoning by referring to a force and a displacement.

- i. the work done on the gas during process I ( $W_I$ )

For the volume to increase, but pressure remains the same. The process must be heated by something. Work is negative

- ii. the work done on the gas during process II ( $W_{II}$ )

This is the same thing as process 2, but at a different constant pressure, so my answer is the same as above (c-i)

- d. During process I, is the heat transfer to the gas positive, negative, or zero? Explain.

Negative, this piston would be moving outward.

# Tutorials In Introductory Physics Homework Answer Key

**M Carnoy**



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### **Table of Contents Tutorials In Introductory Physics Homework Answer Key**

1. Understanding the eBook Tutorials In Introductory Physics Homework Answer Key
  - The Rise of Digital Reading Tutorials In Introductory Physics Homework Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Tutorials In Introductory Physics Homework Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Tutorials In Introductory Physics Homework Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Tutorials In Introductory Physics Homework Answer Key
  - Personalized Recommendations
  - Tutorials In Introductory Physics Homework Answer Key User Reviews and Ratings

- Tutorials In Introductory Physics Homework Answer Key and Bestseller Lists
- 5. Accessing Tutorials In Introductory Physics Homework Answer Key Free and Paid eBooks
  - Tutorials In Introductory Physics Homework Answer Key Public Domain eBooks
  - Tutorials In Introductory Physics Homework Answer Key eBook Subscription Services
  - Tutorials In Introductory Physics Homework Answer Key Budget-Friendly Options
- 6. Navigating Tutorials In Introductory Physics Homework Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Tutorials In Introductory Physics Homework Answer Key Compatibility with Devices
  - Tutorials In Introductory Physics Homework Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Tutorials In Introductory Physics Homework Answer Key
  - Highlighting and Note-Taking Tutorials In Introductory Physics Homework Answer Key
  - Interactive Elements Tutorials In Introductory Physics Homework Answer Key
- 8. Staying Engaged with Tutorials In Introductory Physics Homework Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Tutorials In Introductory Physics Homework Answer Key
- 9. Balancing eBooks and Physical Books Tutorials In Introductory Physics Homework Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Tutorials In Introductory Physics Homework Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Tutorials In Introductory Physics Homework Answer Key
  - Setting Reading Goals Tutorials In Introductory Physics Homework Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Tutorials In Introductory Physics Homework Answer Key
  - Fact-Checking eBook Content of Tutorials In Introductory Physics Homework Answer Key
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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