

Chapter 4 Atomic Structure

Section 4.2 The Structure of an Atom**(pages 108–112)**

This section compares the properties of three subatomic particles. It also discusses atomic numbers, mass numbers, and isotopes.

Reading Strategy (page 108)

Monitoring Your Understanding Before you read, list in the table shown what you know about atoms and what you would like to learn. After you read, list what you have learned. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

What I Know About Atoms	What I Would Like to Learn	What I Have Learned
Most students will know that atoms are the “building blocks” of matter, and some may know that atoms contain subatomic particles.	Based on the title of the section, students may say that they want to learn more about the structure of atoms.	

Properties of Subatomic Particles (pages 108–109)

- What are three subatomic particles?
a. Protons b. Electrons c. Neutrons
- Circle the letter that identifies a subatomic particle with a positive charge.
a. nucleus **(b)** proton
c. neutron d. electron
- Why did Chadwick conclude that the particles produced by his experiment were neutral in charge? A charged object did not deflect the paths of the particles.

Comparing Subatomic Particles (pages 109–110)

- Circle the letters of properties that vary among subatomic particles.
a. color **(b)** mass
(c) charge **(d)** location in the atom
- Circle the letter of the expression that accurately compares the masses of neutrons and protons.
(a) mass of 1 neutron = mass of 1 proton
b. mass of 2000 neutrons = mass of 1 proton
c. mass of 1 electron = mass of 1 proton
d. mass of 1 neutron = mass of 1 electron

Chapter 4 Atomic Structure Answers

Rex M. Heyworth



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Table of Contents Chapter 4 Atomic Structure Answers

1. Understanding the eBook Chapter 4 Atomic Structure Answers
 - The Rise of Digital Reading Chapter 4 Atomic Structure Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 4 Atomic Structure Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 4 Atomic Structure Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 4 Atomic Structure Answers
 - Personalized Recommendations
 - Chapter 4 Atomic Structure Answers User Reviews and Ratings
 - Chapter 4 Atomic Structure Answers and Bestseller Lists
5. Accessing Chapter 4 Atomic Structure Answers Free and Paid eBooks
 - Chapter 4 Atomic Structure Answers Public Domain eBooks
 - Chapter 4 Atomic Structure Answers eBook Subscription Services
 - Chapter 4 Atomic Structure Answers Budget-Friendly Options

6. Navigating Chapter 4 Atomic Structure Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 4 Atomic Structure Answers Compatibility with Devices
 - Chapter 4 Atomic Structure Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 4 Atomic Structure Answers
 - Highlighting and Note-Taking Chapter 4 Atomic Structure Answers
 - Interactive Elements Chapter 4 Atomic Structure Answers
8. Staying Engaged with Chapter 4 Atomic Structure Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 4 Atomic Structure Answers
9. Balancing eBooks and Physical Books Chapter 4 Atomic Structure Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 4 Atomic Structure Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 4 Atomic Structure Answers
 - Setting Reading Goals Chapter 4 Atomic Structure Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 4 Atomic Structure Answers
 - Fact-Checking eBook Content of Chapter 4 Atomic Structure Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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