

## Chapter 6 Thermochemistry

### Homework Problems

- The enthalpy change for the fermentation (without oxygen,  $O_2$  gas) reaction of glucose to alcohol and carbon dioxide is  $-67$  kJ.
  - Write a chemical equation for the fermentation reaction and show how energy is involved in the reaction.
  - Is the reaction exothermic or endothermic?
  - Is energy, in the form of heat, absorbed or evolved as the reaction occurs?
- The specific heat capacity of graphite is  $0.71$  J/°C·g. Calculate the energy required to:
  - Raise the temperature of  $1.0$  mol graphite by  $1.0$ °C.
  - Raise the temperature of  $850$  g graphite by  $150$ °C.
  - Raise the temperature of  $75$  kg graphite from  $294$  K to  $348$  K.
- In a coffee-cup calorimeter,  $100.0$  mL of  $1.0$  M NaOH and  $100.0$  mL of  $1.0$  M HCl are mixed. Both solutions were originally at  $24.6$ °C. After the reaction, the final temperature is  $31.3$ °C. Assume that all the solutions have a density of  $1.0$  g/mL and a specific heat capacity of  $4.18$  J/g°C. Calculate the enthalpy change for the neutralization of HCl by NaOH. Assume that no heat is lost to the surroundings or to the calorimeter.
- Camphor ( $C_{15}H_{26}O$ ) has an energy of combustion of  $-5903.6$  kJ/mol. A sample of camphor with a mass of  $0.1204$  g is burned in a bomb calorimeter, the temperature increases by  $2.28$ °C. Calculate the heat capacity of the calorimeter.  
A  $0.1964$  g sample of quinone ( $C_6H_4O_2$ ) is burned in a bomb calorimeter that has a heat capacity of  $1.56$  kJ/°C. The temperature of the calorimeter increases by  $3.2$ °C. Calculate the energy of combustion of quinone per gram and per mole.

# Chapter 6 Thermochemistry Homework Problems

**Williams & Wilkins Review**



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and equipment Practice drills at the end of each content review chapter plus step by step walk throughs of sample exam questions

Solutions Manual to Accompany General Chemistry with Qualitative Analysis, Second Edition Ralph H. Petrucci, Robert K. Wismer, 1987

**Chemistry** Steven S. Zumdahl, Kenneth C. Brooks, 1986

Ebook: Chemistry: The Molecular Nature of Matter and Change Silberberg, 2015-01-16

Ebook Chemistry The Molecular Nature of Matter and Change **CHEMISTRY SILBERBERG**, 2003

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Schaum's Outline of Theory and Problems of College Chemistry Jerome Laib Rosenberg, Lawrence M. Epstein, 1990

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*Engineering News* ,1906      General Chemistry Key Concise Review Guide Sixth Edition Darrell D. Ebbing,1999

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## **Table of Contents Chapter 6 Thermochemistry Homework Problems**

1. Understanding the eBook Chapter 6 Thermochemistry Homework Problems
  - The Rise of Digital Reading Chapter 6 Thermochemistry Homework Problems
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Thermochemistry Homework Problems
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 6 Thermochemistry Homework Problems
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Thermochemistry Homework Problems
  - Personalized Recommendations
  - Chapter 6 Thermochemistry Homework Problems User Reviews and Ratings
  - Chapter 6 Thermochemistry Homework Problems and Bestseller Lists
5. Accessing Chapter 6 Thermochemistry Homework Problems Free and Paid eBooks

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  - Chapter 6 Thermochemistry Homework Problems Budget-Friendly Options
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    - ePub, PDF, MOBI, and More
    - Chapter 6 Thermochemistry Homework Problems Compatibility with Devices
    - Chapter 6 Thermochemistry Homework Problems Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Chapter 6 Thermochemistry Homework Problems
    - Highlighting and Note-Taking Chapter 6 Thermochemistry Homework Problems
    - Interactive Elements Chapter 6 Thermochemistry Homework Problems
  8. Staying Engaged with Chapter 6 Thermochemistry Homework Problems
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Chapter 6 Thermochemistry Homework Problems
  9. Balancing eBooks and Physical Books Chapter 6 Thermochemistry Homework Problems
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Chapter 6 Thermochemistry Homework Problems
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Chapter 6 Thermochemistry Homework Problems
    - Setting Reading Goals Chapter 6 Thermochemistry Homework Problems
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Chapter 6 Thermochemistry Homework Problems
    - Fact-Checking eBook Content of Chapter 6 Thermochemistry Homework Problems
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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