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chapter - Chemical Kinetics - 01.

Admit
30/06/2023

i) Instantaneous Rate of Reaction.

- (a) It occurs within a short span of time.
- (b) It cannot be calculated for multistep ~~elementary~~ reaction.
- (c) It can be calculated for elementary reaction.

Average Rate of Reaction.

- (a) It occurs during a long interval of time.
- (b) It can be calculated for multistep reaction.
- (c) It cannot be calculated for elementary reaction.

ii) If the unit of rate constant is s^{-1} , the reaction is of first order.

iii) No, the molecularity can never be zero or a fractional number as it shows the number of reactants taking part in a reaction can never be zero.

iv)



Rate of rxn (M) = $k[R]^n$ \rightarrow order of rxn. Reaction involves two reactants.

? $x = k[R]^n$ — (1)

$8x = k[2R]^n$ — (2)

Dividing eq (2) by (1);

$$\frac{8x}{x} = \frac{k[2R]^n}{k[R]^n}$$

or, $8 = 2^n \cdot \frac{R^n}{R^n}$

Overall order

= (a+b)

= (1+2) = 3

or, $8 = 2^n$

or, $2^3 = 2^n$

$\therefore n = 3$

The order of reaction is 3.

(1) $y = k[A]^a[B]^b$

(2) $2y = k[2A]^a[2B]^b$

(3) $8y = k[2A]^a[2B]^b$

(2) \div (1) \rightarrow

$$\frac{2y}{y} = \frac{k[2A]^a[2B]^b}{k[A]^a[B]^b}$$

$\sim 2^1 = 2^a \cdot 2^b$ $\sim a=1$

(3) \div (2) \rightarrow

$2^b = 4 = 2^2 \sim b=2$

(v) The rate constant for the first order reaction is $60 s^{-1}$.

let M be the molar mass.
a be the initial concentration.

$k = 60/s$

final concentration = $\frac{M}{16} \times a$

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Table of Contents Chemsheets Kinetics 1 Answers

1. Understanding the eBook Chemsheets Kinetics 1 Answers
 - The Rise of Digital Reading Chemsheets Kinetics 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemsheets Kinetics 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemsheets Kinetics 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemsheets Kinetics 1 Answers

- Personalized Recommendations
 - Chemsheets Kinetics 1 Answers User Reviews and Ratings
 - Chemsheets Kinetics 1 Answers and Bestseller Lists
5. Accessing Chemsheets Kinetics 1 Answers Free and Paid eBooks
 - Chemsheets Kinetics 1 Answers Public Domain eBooks
 - Chemsheets Kinetics 1 Answers eBook Subscription Services
 - Chemsheets Kinetics 1 Answers Budget-Friendly Options
 6. Navigating Chemsheets Kinetics 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
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 - Chemsheets Kinetics 1 Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemsheets Kinetics 1 Answers
 - Highlighting and Note-Taking Chemsheets Kinetics 1 Answers
 - Interactive Elements Chemsheets Kinetics 1 Answers
 8. Staying Engaged with Chemsheets Kinetics 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemsheets Kinetics 1 Answers
 9. Balancing eBooks and Physical Books Chemsheets Kinetics 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemsheets Kinetics 1 Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemsheets Kinetics 1 Answers
 - Setting Reading Goals Chemsheets Kinetics 1 Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemsheets Kinetics 1 Answers

-
- Fact-Checking eBook Content of Chemsheets Kinetics 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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