

## Calorimetry Lab (aka: "Burn Lab")

https://www.youtube.com/watch?v=...

### Introduction:

By measuring the heat that a chemical energy source or molecular bond is released as heat and light. The heat can be measured in some solid calories (7 calories equal 1000 Joules). Heat energy is used in the food industry. The more calories a food contains, the more heat energy it will provide. Foods high in calories will release large amounts of energy. One gram of a particular substance will release a certain amount of heat.

As calories is the amount of heat (energy) required to increase the temperature of 1 gram of water by 1°C. This process is the basis of the technique of calorimetry. In thermal investigations, you will use the method of calorimetry to approximate the amount of energy contained in a particular food unit or other food item. This heat goes off from the calorimeter as indicated by water that is evaporated above the burning food item. This method approximates the amount of heat given off by the food.

### Purpose:

1. To determine how many calories of energy burning food items contain & compare to real number as indicated by the manufacturers?
2. Determine how this experiment can be improved to produce more accurate results.

### Materials:

- |                        |            |                         |                  |
|------------------------|------------|-------------------------|------------------|
| a. Test tubes & holder | b. Baggies | c. Thermometer or scale | d. matches       |
| e. Copper/calorimeter  | f. water   | g. Food items           | h. 1 Paper clips |

### Procedure:

1. Prepare your calorimeter above the evaporating dish using 1 paper clip. The paper calorimeter (the alcohol thermometer "burned" method) will be used to measure the heat of the food.
2. Measure the mass of the test tube with the evaporating dish paper clip set up above the burner. Record in data table.
3. Weigh 5.00 grams of water (mass) in a beaker and pour it into the test tube. Record the mass of  $H_2O$  in the data table (mass = 5.00 g, vol  $H_2O$  = 5.00 ml).
4. Place the clip in the evaporating dish on the base of the ring stand. Light the test tube holder as that it will be directly above the clip. Check the thermometer for the initial and record.
5. Measure and record the initial temperature of the water.
6. Immediately move the calorimeter to be over the test tube. Use a match to ignite the bottom of the test tube and then put the burning test tube under the test tube. The calorimeter is held up above the burner from the burning test tube ring stand base the calorimeter is held up above the burner.
7. Light the test tube and let the calorimeter be over the test tube until the test tube burns. Measure the highest temperature of the water and record it in the data table. The light clip is it goes off.
8. Measure the final mass of the "clip" set in the evaporating dish of the burning (burned) test.
9. Repeat steps 1-7 using a different food item.

# Calorimetry Lab Aka Burn Answers

**Joacim Rocklöv**



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