

## Balancing Equations Worksheet

- 1) \_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_  $\text{KOH}$   $\rightarrow$  \_\_\_\_  $\text{NaOH}$  + \_\_\_\_  $\text{K}_3\text{PO}_4$
- 2) \_\_\_\_  $\text{MgF}_2$  + \_\_\_\_  $\text{Li}_2\text{CO}_3$   $\rightarrow$  \_\_\_\_  $\text{MgCO}_3$  + \_\_\_\_  $\text{LiF}$
- 3) \_\_\_\_  $\text{P}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{P}_2\text{O}_3$
- 4) \_\_\_\_  $\text{RbNO}_3$  + \_\_\_\_  $\text{BeF}_2$   $\rightarrow$  \_\_\_\_  $\text{Be}(\text{NO}_3)_2$  + \_\_\_\_  $\text{RbF}$
- 5) \_\_\_\_  $\text{AgNO}_3$  + \_\_\_\_  $\text{Cu}$   $\rightarrow$  \_\_\_\_  $\text{Cu}(\text{NO}_3)_2$  + \_\_\_\_  $\text{Ag}$
- 6) \_\_\_\_  $\text{CF}_4$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{CBr}_4$  + \_\_\_\_  $\text{F}_2$
- 7) \_\_\_\_  $\text{HCN}$  + \_\_\_\_  $\text{CuSO}_4$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{SO}_4$  + \_\_\_\_  $\text{Cu}(\text{CN})_2$
- 8) \_\_\_\_  $\text{GaF}_3$  + \_\_\_\_  $\text{Cs}$   $\rightarrow$  \_\_\_\_  $\text{CsF}$  + \_\_\_\_  $\text{Ga}$
- 9) \_\_\_\_  $\text{BaS}$  + \_\_\_\_  $\text{PtF}_2$   $\rightarrow$  \_\_\_\_  $\text{BaF}_2$  + \_\_\_\_  $\text{PtS}$
- 10) \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_\_  $\text{NH}_3$
- 11) \_\_\_\_  $\text{NaF}$  + \_\_\_\_  $\text{Br}_2$   $\rightarrow$  \_\_\_\_  $\text{NaBr}$  + \_\_\_\_  $\text{F}_2$
- 12) \_\_\_\_  $\text{Pb}(\text{OH})_2$  + \_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_  $\text{H}_2\text{O}$  + \_\_\_\_  $\text{PbCl}_2$
- 13) \_\_\_\_  $\text{AlBr}_3$  + \_\_\_\_  $\text{K}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_  $\text{KBr}$  + \_\_\_\_  $\text{Al}_2(\text{SO}_4)_3$
- 14) \_\_\_\_  $\text{CH}_4$  + \_\_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_\_  $\text{CO}_2$  + \_\_\_\_  $\text{H}_2\text{O}$
- 15) \_\_\_\_  $\text{Na}_3\text{PO}_4$  + \_\_\_\_  $\text{CaCl}_2$   $\rightarrow$  \_\_\_\_  $\text{NaCl}$  + \_\_\_\_  $\text{Ca}_3(\text{PO}_4)_2$
- 16) \_\_\_\_  $\text{K}$  + \_\_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_\_  $\text{KCl}$
- 17) \_\_\_\_  $\text{Al}$  + \_\_\_\_  $\text{HCl}$   $\rightarrow$  \_\_\_\_  $\text{H}_2$  + \_\_\_\_  $\text{AlCl}_3$
- 18) \_\_\_\_  $\text{N}_2$  + \_\_\_\_  $\text{F}_2$   $\rightarrow$  \_\_\_\_  $\text{NF}_3$
- 19) \_\_\_\_  $\text{SO}_2$  + \_\_\_\_  $\text{Li}_2\text{Se}$   $\rightarrow$  \_\_\_\_  $\text{SSe}_2$  + \_\_\_\_  $\text{Li}_2\text{O}$
- 20) \_\_\_\_  $\text{NH}_3$  + \_\_\_\_  $\text{H}_2\text{SO}_4$   $\rightarrow$  \_\_\_\_  $(\text{NH}_4)_2\text{SO}_4$

# Balancing Equation Practice Packet Answer Sheet

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## **Balancing Equation Practice Packet Answer Sheet:**

**The Practice of Chemistry Study Guide & Solutions Manual** Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises *CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE* Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text **Chemistry, Student Study Guide** James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement *Study Guide to Accompany Basics for Chemistry* Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students *Educart CBSE Sample Paper Class 12 2025-26 - Chemistry (Based on latest CBSE Sample Papers Released on 30-07-25) For 2026 Board Exam* Educart, 2025-10-09 This book

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third edition continues to innovate by providing students with an integrated and modern approach to the subject. The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors. Also of benefit is the just in time presentation of key content only providing details once they are needed. While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium. The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy throughout complete with text icons. The whole is backed by a range of supplements including a new illustration program, a tutorial CD, interactive learningware, an extensive Web CT component, an instructor's resource CD, and a solution CD.

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Julie R. Frentrop, 2002-06. Contains a brief overview of every chapter, review of skills, self tests, and the answers and detailed solutions to all odd numbered end of chapter problems in the text book.

Study Guide and Problems Book for Biochemistry, Garrett and Grisham  
David Karl Jemiolo, 1995. Continuing Garrett and Grisham's innovative conceptual and organizing framework. Essential Questions BIOCHEMISTRY guides students through course concepts in a way that reveals the beauty and usefulness of biochemistry in the everyday world. Streamlined for increased clarity and readability, this edition also includes new photos and illustrations that show the subject matter consistently throughout the text. New end of chapter problems, MCAT practice questions, and the unparalleled text media integration with the power of CengageNOW round out this exceptional package, giving you the tools you need to both master course concepts and develop critical problem solving skills you can draw upon long after the course ends.

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