

**Introduction:**

When you bake/cook something, we use a specific amount of each ingredient. Imagine if you made a batch of cookies and used way too many eggs, or not enough sugar. YUCK! In chemistry, reactions proceed with very specific recipes. The study of these recipes is stoichiometry. When the reactants are present in the correct amounts, the reaction makes products. We are going to look at what happens when there are more or less of some of the reactants present.

**Procedure:** PHET Simulations → Play with the Sims → Chemistry → Reactions, Products, and Leftovers **OPEN IN SIMS** & [https://phet.colorado.edu/en/simulations/versions/products\\_and\\_leftovers](https://phet.colorado.edu/en/simulations/versions/products_and_leftovers) **products\_and\_leftovers** [https://phet.colorado.edu/en/simulations/versions/products\\_and\\_leftovers](https://phet.colorado.edu/en/simulations/versions/products_and_leftovers)

**Part 1: Making Sandwiches**



1. A cheese sandwich requires 2 slices of bread to 1 slice of cheese. If there is exactly twice as many slices of bread as cheese, then there will be no leftovers. What happens if you have two slices of bread and one slice of cheese? You have a slice of cheese left over!
2. Use the simulation to play around with this two-reactant synthesis reaction. Take some time and familiarize yourself with the simulation buttons.
3. At the bottom of the screen, change the coefficients of the reactants to 2:1.



4. Notice that there is now one sandwich with no leftovers!
5. Complete situations B through D. A has been done for you as an example.

Situation	Bread Used	Cheese Used	Sandwiches Made	Excess Bread	Excess Cheese
A	4 slices	2 slices	2 sandwiches	0 bread	0 cheese
B	2 slices	3 slices			
C			2 sandwiches	1 slice	0 slices
D	6 slices		3 sandwiches		4 slices

6. Look at situation A. The cheese in this case is said to be in excess. The other reactant (bread) has no leftovers. The bread in this case is the limiting reactant.
7. In situations B, C, and D, determine the limiting reactant: B: \_\_\_\_\_ C: \_\_\_\_\_ D: \_\_\_\_\_



**Part 2: Real Chemical Reactions**



8. Select "Molecules" towards the bottom of your screen. We can now work with real chemical reactions, one that creates a very important molecule!
9. What is the lowest molecular ratio for the reaction of hydrogen and oxygen to produce water? That is, balance the following equation:



10. Complete the table below while making water H<sub>2</sub>O from hydrogen H<sub>2</sub> and oxygen O<sub>2</sub>.

Hydrogen Molecules H <sub>2</sub>	Oxygen Molecules O <sub>2</sub>	Water Molecules H <sub>2</sub> O	Excess H <sub>2</sub>	Excess O <sub>2</sub>
4 molecules	4 molecules			
7 molecules	6 molecules			
		4 molecules	0 molecules	0 molecules
6 molecules	8 molecules			
		4 molecules	1 molecules	0 molecules

Stop for check point. Teacher signature for this page: \_\_\_\_\_

# Basic Stoichiometry Phet Lab Homework Exercises Answers

**Minjie Lin**



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## **Table of Contents Basic Stoichiometry Phet Lab Homework Exercises Answers**

1. Understanding the eBook Basic Stoichiometry Phet Lab Homework Exercises Answers
  - The Rise of Digital Reading Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Basic Stoichiometry Phet Lab Homework Exercises Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Personalized Recommendations
  - Basic Stoichiometry Phet Lab Homework Exercises Answers User Reviews and Ratings
  - Basic Stoichiometry Phet Lab Homework Exercises Answers and Bestseller Lists
5. Accessing Basic Stoichiometry Phet Lab Homework Exercises Answers Free and Paid eBooks
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Public Domain eBooks
  - Basic Stoichiometry Phet Lab Homework Exercises Answers eBook Subscription Services
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Budget-Friendly Options
6. Navigating Basic Stoichiometry Phet Lab Homework Exercises Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Compatibility with Devices
  - Basic Stoichiometry Phet Lab Homework Exercises Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Highlighting and Note-Taking Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Interactive Elements Basic Stoichiometry Phet Lab Homework Exercises Answers

8. Staying Engaged with Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Basic Stoichiometry Phet Lab Homework Exercises Answers
9. Balancing eBooks and Physical Books Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Basic Stoichiometry Phet Lab Homework Exercises Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Setting Reading Goals Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Fact-Checking eBook Content of Basic Stoichiometry Phet Lab Homework Exercises Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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