

8.2 – THE CHEMICAL EARTH

1 – The living and non-living components of the Earth contain mixtures

1.1 Construct word and balanced formulae equations of chemical reactions as they are encountered

- To make **word equations**, write each chemical formula, in words. Then, add in pluses and arrows.



- To create **balanced formulae equations**, write down **chemical symbols** (see 1.7), their **state of matter** and **balance** the equation.

Magnesium + Oxygen \rightarrow Magnesium Oxide	The original word equation.
$\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$	There are not enough Oxygen atoms on the right.
$\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$	Double the whole molecule , not just magnesium. There are not enough magnesium atoms now.
$2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$	Add a 2 in front of magnesium on the left. Now, LHS = RHS
$2\text{Mg}_{(s)} + \text{O}_{2(g)} \rightarrow 2\text{MgO}_{(s)}$	Place the states of matter for each substance.

- Combustion – **Fuel + Oxygen \rightarrow Carbon dioxide + Water**, example: $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
- Corrosion – **Metal + Oxygen \rightarrow Metal Oxide**, example: $2\text{Cu} + \text{O}_2 \rightarrow 2\text{CuO}$
- Respiration – **Glucose + Oxygen \rightarrow Carbon Dioxide + Water**, example: $\text{C}_6\text{H}_{12}\text{O}_6 + 6\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$
- Precipitation – **Soluble salt + Soluble salt \rightarrow Insoluble salt + Soluble salt**
 - o Example: $\text{CuSO}_4 + 2\text{NaOH} \rightarrow \text{Cu}(\text{OH})_2 + \text{Na}_2\text{SO}_4$ ($\text{Cu}(\text{OH})_2$ is insoluble)
- Neutralisation – **Acid + Base \rightarrow Salt + Water**, example: $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
- Acids on metals – **Metal + Acid \rightarrow Salt + Hydrogen**, example: $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$
- Acid on metal oxide – **Metal Oxide + Acid \rightarrow Salt + Water**, example: $\text{MgO} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$
- Acids on carbonates – **Carbonate + Acid \rightarrow Salt + Water + Carbon Dioxide**
 - o Example: $\text{CaCO}_3 + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$
- Decomposition – **Molecule \rightarrow Molecule part + Molecule part**, example: $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$

1.2 Identify the difference between elements, compounds and mixtures in terms of particle theory

- Particle theory helps us **visualise** differences between elements, compounds and mixtures.
- **Elements** are basic building blocks, having own unique symbol and physical and chemical properties.
- **Compounds** are **two or more particles** chemically bonded.
- A **mixture** is composed of a mix of **elements and compounds**, e.g. a solution.
 - o Can be separated using simple separation techniques, does not have fixed properties.

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