

17. SOLUBILITY PRODUCT CONSTANT

GOALS

1. Become aware of solubility product constants
2. Estimate a solubility constant
3. Observe that a solubility constant is a special case of an equilibrium constant

BACKGROUND

Ionic compounds, such as salts, dissociate to form ions when they dissolve in water. An example is sodium chloride:



Where *aq* is a solution and *s* is a solid.

Solid *NaCl* dissociates to become *Na⁺* and *Cl⁻* ions. The *Na⁺* cations are attracted to the negative "ends" of water molecules, and the *Cl⁻* anions are attracted to the positive "ends" of water molecules.

In other compounds the attraction between the cations and the anions is stronger and there is little dissociation. These are sparingly or slightly soluble salts. An example is silver chloride, *AgCl*. A beaker of this solution often has solid *AgCl* crystals in the bottom. The crystals are in equilibrium with the *Ag⁺* and *Cl⁻* ions that are in solution:



In this reaction the rate of *AgCl* dissolving is equal to the rate of *Ag⁺* and *Cl⁻* ions combining to form solid *AgCl*. In other words, the reaction is going both ways at the same time and at the same rate. It is said to be in equilibrium. The equilibrium constant is:



Solubility Product Constant Lab Activity

Minjie Lin



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Exploring General Chemistry in the Laboratory Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science

Illustrated Guide to Home Chemistry Experiments Robert Bruce Thompson, 2008-04-29 For students DIY hobbyists and science buffs who can no longer get real chemistry sets this one of a kind guide explains how to set up and use a home chemistry lab with step by step instructions for conducting experiments in basic chemistry not just to make pretty colors and stinky smells but to learn how to do real lab work Purify alcohol by distillation Produce hydrogen and oxygen gas by electrolysis Smelt metallic copper from copper ore you make yourself Analyze the makeup of seawater bone and other common substances Synthesize oil of wintergreen from aspirin and rayon fiber from paper Perform forensics tests for fingerprints blood drugs and poisons and much more From the 1930s through the 1970s chemistry sets were among the most popular Christmas gifts selling in the millions But two decades ago real chemistry sets began to disappear as manufacturers and retailers became concerned about liability em The Illustrated Guide to Home Chemistry Experiments steps up to the plate with lessons on how to equip your home chemistry lab master laboratory skills and work safely in your lab The bulk of this book consists of 17 hands on chapters that include multiple laboratory sessions on the following topics Separating Mixtures Solubility and Solutions Colligative Properties of Solutions Introduction to Chemical Reactions Stoichiometry Reduction Oxidation Redox Reactions Acid Base Chemistry Chemical Kinetics Chemical Equilibrium and Le Chatelier s Principle Gas Chemistry Thermochemistry and Calorimetry Electrochemistry Photochemistry Colloids and Suspensions Qualitative Analysis Quantitative Analysis Synthesis of Useful Compounds Forensic Chemistry With plenty of full color illustrations and photos Illustrated Guide to Home Chemistry Experiments offers introductory level sessions suitable for a middle school or first year high school chemistry laboratory course and more advanced sessions suitable for students who intend to take the College Board Advanced Placement AP Chemistry exam A student who completes all of the laboratories in this book will have done the equivalent of two full years of high school chemistry lab work or a first year college general chemistry laboratory course This hands on introduction to real chemistry using real equipment real chemicals and real quantitative experiments is ideal for the many thousands of young people and adults who want to experience the magic of chemistry

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Laboratory James M. Postma, Julian L. Roberts, J. Leland Hollenberg, 2009-08-07 For nearly 40 years Chemistry in the Laboratory has been meeting the needs of teachers and students This new edition builds on that legacy while addressing cutting edge trends in the chemistry laboratory including forensic chemistry and environmental and green chemistry As always the new edition of Chemistry in the Laboratory offers precise easy to follow instructions helpful illustrations and an emphasis throughout on laboratory safety Again throughout a Consider This feature encourages users to expand the principles of the experiment into interesting applications open ended experiments or unexplored corners Most experiments in the manual can be completed in one lab session but some can be linked or extended for a multi lab project **TID.** ,1956

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