

The density of 3 M solution of NaCl is  $1.25 \text{ g mL}^{-1}$ . Calculate the molality of the solution.

### **Solution**

$$M = 3 \text{ mol L}^{-1}$$

Mass of NaCl

$$\text{in 1 L solution} = 3 \times 58.5 = 175.5 \text{ g}$$

Mass of

$$1 \text{ L solution} = 1000 \times 1.25 = 1250 \text{ g}$$

(since density =  $1.25 \text{ g mL}^{-1}$ )

$$\begin{aligned} \text{Mass of water in solution} &= 1250 - 175.5 \\ &= 1074.5 \text{ g} \end{aligned}$$

$$\text{Molality} = \frac{\text{No. of moles of solute}}{\text{Mass of solvent in kg}}$$

$$= \frac{3 \text{ mol}}{1.0745 \text{ kg}} = 2.79 \text{ m}$$

# Solution Of Kachot

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