

# Bioprocess Engineering

## Mole Concept

$$\text{mole} = \frac{\text{Given mass (kg or gm)}}{\text{molar mass}}$$

$$\text{mole} = \frac{\text{No. of substance}}{N_A (6.02 \times 10^{23})}$$

$$\rightarrow 1 \text{ atom} = \left(\frac{1}{12}\right)^{\text{th}} \text{ of mass of C-12 atom}$$

$$\rightarrow 1 \text{ amu} = 1.66 \times 10^{-24} \text{ gm} \\ \text{or } 1.66 \times 10^{-27} \text{ kg}$$

$$\rightarrow 1 \text{ mole} = 6.02 \times 10^{23} \text{ substance}$$

$$\text{charge of 1 mole of } e^- = (1.66 \times 10^{-19}) \times (6.02 \times 10^{23}) \\ = 96500 \text{ coulomb} = 1 \text{ Farad.}$$

$$\text{molarity (M)} = \frac{\text{moles of solute}}{\text{volume of sol}^n (\text{lit})} \quad \text{unit (m/lit)}$$

# Bioprocess Engineering Notes By Shular

**M Tight**



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