

Chemistry Worksheets Class 11 on Chapter 8 Redox Reactions - Set 1

Q-1: Which of the following statements about redox reactions is correct?

- a) A nonmetal with a higher reduction potential is a more potent reducing agent.
- b) A metal with a lower oxidation potential is a more potent oxidising agent.
- c) The oxygen oxidation state in O_2 is -1.
- d) All of the above

Q-2: Which of the following can behave as a strong reducing agent?

- a) $HClO_4$
- b) $SnCl_2$
- c) SO_2
- d) All of the above

Q-3: The pair of compounds in which both metals are in the highest oxidation state is

- a) $FeSO_4$, CuS_2
- b) CrO_2Cl_2 , MnO_4^-
- c) TiO_2 , MnO_2
- d) $[Co(CN)_6]^{2-}$, MnO_2

Q-4: The number of electrons involved when 1 mole of H_2O_2 decomposes as

- $H_2O_2 \rightarrow H_2O + O_2$
- a) 1
 - b) 6
 - c) 8
 - d) 4

Q-5: Zinc iodide is formed during the reaction of zinc and iodine. Which of these is being oxidised?

- a) Zinc ions
- b) Zinc atom
- c) Iodide ions
- d) Iodine

Q-6: Can a solution of 1M copper sulphate be stored in a vessel made of nickel metal?

Given that $E^\circ(Ni^{2+}/Ni) = -0.25$ volt and $E^\circ(Cu^{2+}/Cu) = +0.34$ volt.

Q-7: Using the electronic concept, identify the oxidant and reductant in the following redox reactions.

- i) $Zn + 2H^+ \rightarrow Zn^{2+} + H_2$
- ii) $BrO_3^- + F_2 + 2OH^- \rightarrow BrO_4^- + 2F^- + H_2O$

Chapter 11 Chemical Reactions Worksheet

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