

Charges, Fields, and Potential

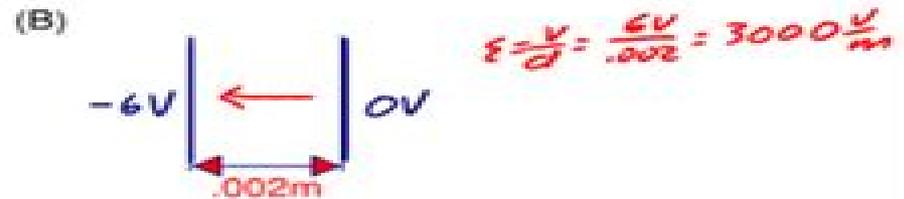
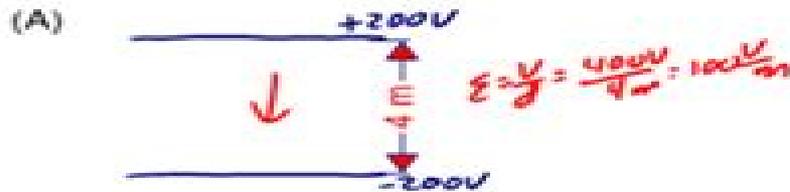
1. A conducting coathanger has a net excess charge of $6.4 \cdot 10^{-19}$ coulombs. The coathanger must:

- (A) have 2 excess electrons
- (B) be missing 2 protons
- (C) have 4 excess electrons
- (D) be missing 4 electrons

2. A metal conducting sphere has a net excess charge of $-4.8 \cdot 10^{-19}$ coulombs. The metal sphere must:

- (A) have 2 excess electrons
- (B) be missing 2 protons
- (C) have 3 excess electrons
- (D) be missing 3 electrons

3. Calculate the electric field for each of the following. Indicate both magnitude and direction.



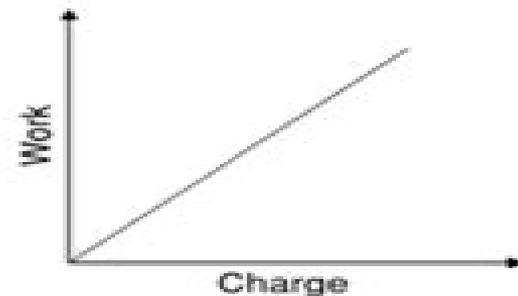
4. Which diagram best represents the electric field around a negatively charged conducting sphere?



5. The graph at right shows the relationship between the work done on a charged body in an electric field and the net charge on the body. What does the slope of this graph represent?

- (A) Power
- (B) Potential difference
- (C) Force
- (D) Electric field intensity

$$\frac{\text{rise}}{\text{run}} = \frac{W}{q} = \frac{8J}{8C} = V$$



6. A metal sphere having an excess of +5 elementary charges has a net electric charge of $8 \cdot 10^{-19}$ C.

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Padhraic Smyth



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Report of the Secretary for Public Instruction ... Queensland. Department of Public Instruction,1901

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Electronics in Our World Gregory J. Nunz,1972 **Science Spectrum** Holt Rinehart & Winston,Holt, Rinehart and Winston Staff,2003-03 **International Aerospace Abstracts** ,1977 *Gramophone Record Review* ,1956

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds $\text{Cl}^- \text{aq}$

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver I^+ ions regardless of the source of these ions these molecular and

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+ \text{aq} + 2\text{F}^- \text{aq} + 2\text{NH}_4^+ \text{aq} + \text{Cr}_2\text{O}_7^{2-} \text{aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{s} + 2\text{NH}_4\text{F} \text{aq}$ because two $\text{NH}_4^+ \text{aq}$

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