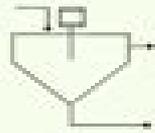
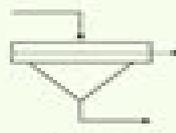


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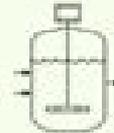
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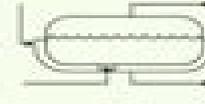
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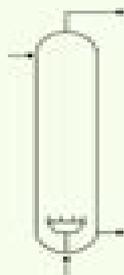
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Vapor



Evaporator



Tower



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Cooling tower



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# Symbols Process Flow Diagram Chemical Engineering

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web balance this redox reaction by the half reaction method by assuming an acidic solution  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} \rightarrow \text{Cr}^{3+} + \text{Fe}^{3+}$  balance the redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^- + \text{Au} + \text{CN}^- + 2\text{e}^- + 2\text{H}^+ + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CNO}^-$  make electrons equal convert to basic solution  $4\text{CN}^- + 2\text{Au} + 2\text{CN}^- + 2\text{e}^-$  multiplied by a factor of 2  $2\text{e}^- + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^-$  the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation  $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$  now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1  $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$  solution 1 balanced half reactions  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$  equalize the electrons  $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$   $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$  multiplied by a factor of 6

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web balance each half reaction in basic solution  $4\text{S} + 6\text{OH}^- + 2\text{NO}_3^- \rightarrow \text{SO}_4^{2-} + 2\text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$  balance each redox reaction in acid solution using the half reaction method  $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{TeO}_3 + 2\text{N}_2\text{O}_4 + \text{O}_2 + \text{Cr}^{3+} + \text{Te}$   $10\text{ReO}_4^- + 11\text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> 2h<sub>2</sub>no<sub>2</sub> c<sub>6</sub>h<sub>6</sub>o<sub>6</sub> 2h<sub>2</sub>o 2no choose 1 answer c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> and no<sub>2</sub> a c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> and no<sub>2</sub> no<sub>2</sub> and c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> b no<sub>2</sub> and c<sub>6</sub>h<sub>8</sub>o<sub>6</sub> h and no<sub>2</sub> c h and

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