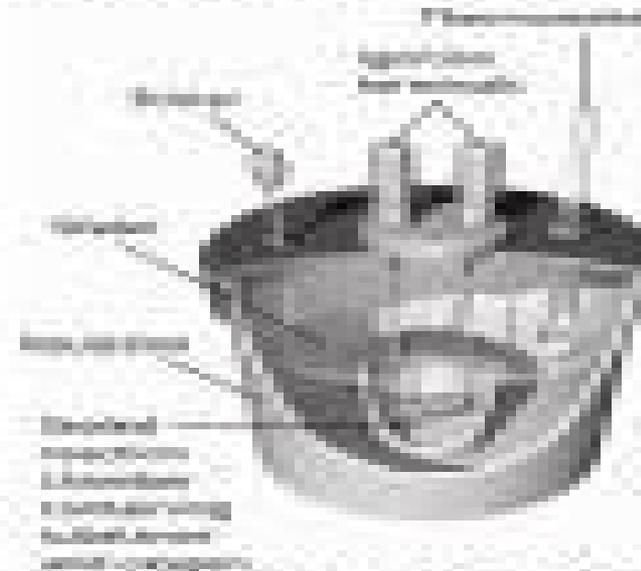


TT44 - Using a Calorimeter



1. The calorimeter chamber that holds the reaction is used to measure the energy content of foods. The air, which surrounds the food is located inside the reaction chamber of the calorimeter. Air is the system which is the surroundings. The system is the reaction chamber and the surroundings are the liquid, water, insulation.
2. When heated food is added to the chamber, the energy which must also be added to measure the temperature. **Energy is lost to the system.**
3. Heat is the product of all the reactions that take place in the reaction chamber. The product is the heat measurement.
4. This is the calorimeter chamber and it is used to keep the heat of the liquid for better measurement.
5. When the food is added to the chamber, the calorimeter measures the heat.
6. During the process of heat in the reaction, the gas inside the chamber, the liquid and insulation are used and the calorimeter measures the temperature.
7. Assuming that no heat escapes from the calorimeter, when a heated substance is added, the amount of heat released by the burning food is the same amount of heat that is absorbed in the system.
8. Since the power released from the reaction is equal to the energy which is absorbed, the heat released is the same as the heat absorbed and for the reaction.

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F Rizvi

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