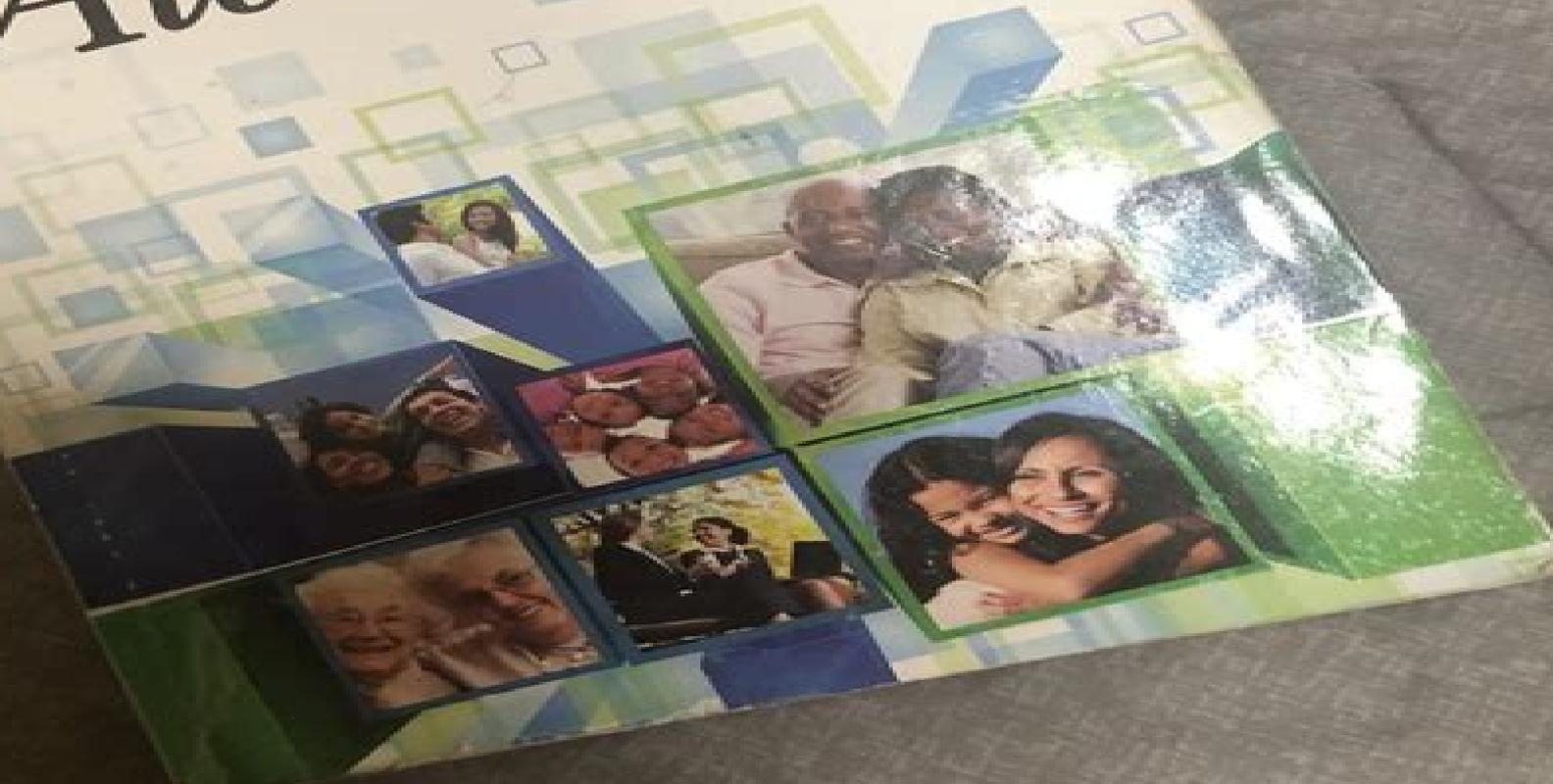


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web a reaction and its experimental rate law are represented below  $2 \text{ICl} + \text{H}_2 \rightarrow 2 \text{HCl} + \text{I}_2$  a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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web reaction rates google classroom  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$  the reaction between  $\text{Mg}$  and  $\text{HCl}$  is represented by the equation above in a kinetics experiment a  $0.080 \text{ g}$  sample of  $\text{Mg}$  ribbon is added to an erlenmeyer flask containing  $50 \text{ ml}$  of  $3.0 \text{ M HCl}$  at  $25^\circ \text{C}$

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