



Determining the Enthalpy Change of a Chemical Reaction

Learning Objectives

Understand the concept of calorimetry, specifically constant-pressure calorimetry.

Measure the temperature difference that occurs during a chemical reaction, using coffee cup calorimetry.

Use the calorimetry data to determine the enthalpy change (ΔH) of the reaction.

$$24.00 \text{ mL} \rightarrow 0.026 \text{ L}$$

$$\left[\left(\frac{4 \text{ mol}}{1 \text{ L}} \right) \cdot 0.026 \text{ L HCl} \left(\frac{1.13 \text{ mol HCl}}{1 \text{ L HCl}} \right) \left(\frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol HCl}} \right) \right] = 0.054932 \text{ mol H}_2\text{O}$$

$$\frac{-51370 \text{ J}}{1 \text{ mol H}_2\text{O}} \left(\frac{0.0549 \text{ mol H}_2\text{O}}{1} \right) = 3149$$

- heat capacity definition: the change in heat required to change the temp. of something by one degree Celsius.



Exploring Chemistry Lab Answers

Julie R. Peller



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