

Vol. of 1 mol of ideal gas at STP/NTP = 22.4 L

उत्स नान एरिया
- 3 same sabse kam significant figure hoga pt. ke bad utne hi hoga ans me.

* Significant fig
• 0 to the left of the 1st non 0 digit are sig. not sig.
eg 0.0012 \Rightarrow 2
• Terminal 0 are not sig. if there is no decimal pt.
eg 1000 \Rightarrow 1 fig.
• 0 at end is sig. if there is decimal pt.
• Exponential term are not considered sig.

अंश - 3/131
axb=c
- add/subst
is least amount
3 same sabse kam fig hoga vo sig. fig ans me hoga

* LAW OF MASS CONSV
• In a reactⁿ \rightarrow actual mass of reactant = mass of prod
Rakme zitna use hua.

* Exceptⁿ \rightarrow Nuclear Rxn (E=mc²)
• PpWedⁿ \rightarrow Balanced Rxn

MOLE CONCEPT

* LAW OF CONST PROPORTⁿ
• Elements in a comp are in fixed ratio by mass.
eg CO₂
12 32g \Rightarrow 3:8
% of C = $\frac{3}{11} \times 100$

Q. cal. the mass of C req to form 132g of CO₂.
 \Rightarrow C + 2O \rightarrow CO₂
12g 32g 44g
44g CO₂ ke lie 12g C
 \therefore 132g ke lie
= $\frac{12}{44} \times 132 = 36\text{g}$

* Law of Multiple Propⁿ
• elements should be same or same fixed mass
eg CO, CO₂ 1:2 ratio
Other eg. N₂O, N₂O₃, N₂O₅
exceptⁿ \rightarrow N₂O₃:1
* Avogadro law

* Same (P, T, R, V) of molecules of gases same (mass eg! no. of)
X₂, X₃, X₄
R : R : R molecules
2 : 3 : 4 atoms [different]

* Gay Lussac's law
CONDINS: +T, P same
• at least 2 gases.
H₂ + Cl₂ \rightarrow 2HCl
1 : 1 : 2
• Vol. R₂ U₂

* DALTON'S ATOMIC THEORY
• Atoms \rightarrow Neither creat nor destroyed

* MOLE CONCEPT
Atomic No. (P) \rightarrow mass no. (P+N)
eg $\frac{A}{Z} X$
• Mass of Proton = $1.67 \times 10^{-27} \text{ kg}$
• Neutron = $1.67 \times 10^{-27} \text{ kg}$
• Electron = $9.1 \times 10^{-31} \text{ kg}$
• MP $\approx 1837 \text{ MeV}$

* Atomic mass unit (amu)
1 amu = $\frac{1}{12}$ Mass of ¹²C
 $1 \text{ amu} = 1.67 \times 10^{-27} \text{ kg}$
eg Mass of ¹²C = 12 amu
* Relative atomic mass
• Unitless
• eg. mass. A.M. of C = 12

* Molecular Mass
• H₂SO₄ = 98 amu
 \rightarrow Mass of 1 molecule of H₂SO₄
* FORMULA MASS: for ionic comp.

* MOLE
Amount of substance.
• 1 mole = No particles (6.022 x 10²³)
1 amu $\leq 1.67 \times 10^{-27} \text{ kg}$
Na = 6.022 x 10²³
1 amu y Na $\leq 1\text{g}$

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AW Rasmussen



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